

**Step 3** From the number of moles of  $\text{Fe}^{3+}$  and  $\text{SCN}^-$  initially present in each mixture, and the number of moles of  $\text{Fe}^{3+}$  and  $\text{SCN}^-$  used up in forming  $\text{FeSCN}^{2+}$ , calculate the number of moles of  $\text{Fe}^{3+}$  and  $\text{SCN}^-$  that remain in each mixture at equilibrium. Use Equation 4. Enter the results in columns 3 and 4 of the table.

**Step 4** Use Equation 3 and the results of Step 3 to find the concentrations of all of the species at equilibrium. The volume of the mixture is 10.00 mL, or 0.0100 liter, in all cases. Enter the values in columns 5 and 7 of the table.

**Step 5** Calculate  $K_c$  for the reaction for each of the mixtures by substituting values for the equilibrium concentrations of  $\text{Fe}^{3+}$ ,  $\text{SCN}^-$ , and  $\text{FeSCN}^{2+}$  in Equation 2.

**Step 6** Calculate the mean value for  $K_c$  and the standard deviation. (See Appendix VIII.)

Mixture	Initial Moles		Equilibrium Moles			Equilibrium Concentration			$K_c$
	$\text{Fe}^{3+}$	$\text{SCN}^-$	$\text{Fe}^{3+}$	$\text{SCN}^-$	$\text{FeSCN}^{2+}$	$[\text{Fe}^{3+}]$	$[\text{SCN}^-]$	$[\text{FeSCN}^{2+}]$	
1	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4}$	$\text{---} \times 10^{-4} \text{ M}$	$\text{---} \times 10^{-4} \text{ M}$	$\text{---} \times 10^{-4} \text{ M}$	$\text{---}$
2	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$
3	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$
4	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$
5	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$	$\text{---}$

Mean value of  $K_c$   $\text{---}$  Standard deviation  $\text{---}$

(continued on following page)

# Experiment 23 Determination Equilibrium Constant

**Jo A. Beran**



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