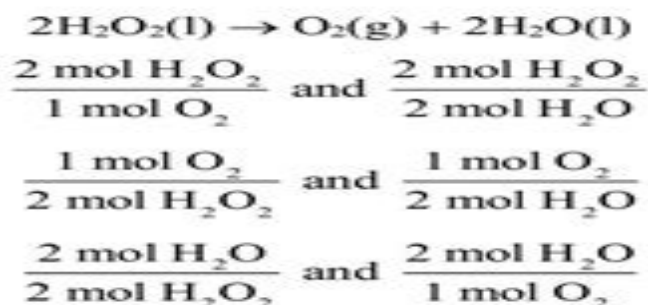


## Chapter 11 (continued)



To determine the number of mole ratios that defines a given chemical reaction, multiply the number of species in the equation by the next lower number. Thus, a chemical reaction with three participating species can be defined by six mole ratios ( $3 \times 2 = 6$ ); a chemical reaction with four species can be defined by 12 mole ratios ( $4 \times 3 = 12$ ).

Why learn to write mole ratios? They are the key to calculations that are based on chemical equations. Using a balanced chemical equation, mole ratios derived from the equation, and a given amount of one of the reactants or products, you can calculate the amount of any other participant in the reaction.

### Practice Problems

2. Determine all the mole ratios for the following balanced chemical equations.
  - a.  $\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{NO}(\text{g})$
  - b.  $4\text{NH}_3(\text{aq}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{l})$
  - c.  $4\text{HCl}(\text{aq}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + 2\text{Cl}_2(\text{g})$

## 11.2 Stoichiometric Calculations

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There are three basic stoichiometric calculations: mole-to-mole conversions, mole-to-mass conversions, and mass-to-mass conversions. All stoichiometric calculations begin with a balanced equation and mole ratios.

► **Stoichiometric mole-to-mole conversion** How can you determine the number of moles of table salt (NaCl) produced from 0.02 moles of chlorine ( $\text{Cl}_2$ )?

# Exercise 11 Stoichiometry 1

**Martin Schmal**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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**Solid-State Electrochemistry** Abdelkader Hammou, Samuel Georges, 2020-02-03 This book features the essential material for any graduate or advanced undergraduate course covering solid state electrochemistry It provides the reader with fundamental course notes and numerous solved exercises making it an invaluable guide and compendium for students of the subject The book places particular emphasis on enhancing the reader's expertise and comprehension of thermodynamics the Kröger-Vink notation the variation in stoichiometry in ionic compounds and of the different types of electrochemical measurements together with their technological applications Containing almost 100 illustrations a glossary and a bibliography the book is particularly useful for Master and PhD students industry engineers university instructors and researchers working with inorganic solids in general

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*Cumulated Index Medicus*, 1975

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*Congressional Record* United States. Congress, 1878

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