

**Criminal
Investigation, a
Practical Handbook
for Magistrates, Police
Officers and Lawyers**

Hans Gross



Criminal Investigation A Practical Handbook For Magistrates Police Officers And Lawyers

DJ Losen



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web limiting reactant theoretical and percentage yields key $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$ If a reaction vessel contains 0.1 mol KO_2 and 0.1 mol H_2O , what is the limiting reactant?

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limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times 1 \text{ mol Al} / 27 \text{ g/mol} = 5.93 \text{ mol Al}$ $5.93 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 / 4 \text{ mol Al} = 2.96 \text{ mol Al}_2\text{O}_3$ $2.96 \text{ mol Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 301.8 \text{ g}$ c determine the percent yield yield actual $\times 100 = 260.0 \text{ g} / 301.8 \text{ g} \times 100 = 86.01\%$ theoretical 302.3 g

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web 7 Na $2\text{C}_2\text{O}_4$ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H_2 are left over 13

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ch₃oh can be produced by the following reaction 2h₂ + co → ch₃oh calculate the theoretical yield of ch₃oh if 68.5 g of co is reacted with 8.6 g of h₂

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