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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} + \text{MnO}_4^{-}$
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redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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convert to basic solution $4\text{cn} + 2\text{au} + 2\text{e}^-$ multiplied by a factor of 2 $2\text{e}^- + 2\text{h}_2\text{O} + \text{O}_2 + 2\text{H}^+ + 3$ the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation Ni(s) → Ni²⁺(aq) + 2e⁻ now that the oxidation half reaction is balanced it tells us that two electrons are produced for

every atom of nickel oxidized but where do those electrons go

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