

ORGO ACIDS & BASES

CHEAT SHEET

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Arrhenius

Acid = Litmus Red
 H^+ in H_2O

Base Litmus Blue
 OH^- in H_2O

Bronsted-Lowry

H^+ donor
 ex. HCl

Accepts H^+
 ex. $NaOH$

Lewis (electrons)

Accepts e^- pair
 ex. $AlCl_3$

e^- pair donor
 ex. Cl^-

Acid - Base Reaction



Equations to recognize

$$K_a = \frac{[H^+][A^-]}{[HA]} \quad pK_a = -\log(K_a) \rightarrow K_a \propto [H^+] \propto \frac{1}{pK_a}$$

Strong Acid

$\uparrow [H^+]$

$\uparrow K_a$

$\downarrow pK_a$

Weak Acid

$\downarrow [H^+]$

$\downarrow K_a$

$\uparrow pK_a$

$$pH = -\log [H^+] \\ pOH = -\log [OH^-]$$

$$pH + pOH = 14 \\ pK_a + pK_b = pK_w = 14$$

pH scale
 $pH \rightarrow$
 $\leftarrow pOH$

$pH = 1$
 Acidic
 $pOH = 14$

$pH = 7$
 Neutral
 $pOH = 7$

Basic
 $pH = 14$
 $pOH = 1$

STRONG ACIDS FORM STABLE CONJUGATE BASES

CARIO

C = charge of acid or conj. base

A = Atom holding charge

R = Resonance

I = Inductive effect

O = orbital / hybridization

See videos for full explanation + examples LEAH4SCI.COM/acidbase

Charge: $+ / \delta^+$ more acidic } When compare species
 $- / \delta^-$ more basic }

Atom: In same period \uparrow eneg \uparrow acidity
 In same group \uparrow size \uparrow acidity



Resonance: \uparrow Res \uparrow charge distribution
 \uparrow acidity (Aromaticity = \uparrow stable)

Inductive Effect: \uparrow e-neg nearby atom
 \uparrow acidity

Orbital / hybridization \uparrow % s \uparrow acidity
 $sp = 50\% > sp^2 = 33\% > sp^3 = 25\%$

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Chemistry Study Guide Acids And Base

Pamela Mills,Amina El-Ashmawy



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and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text

Study Guide and Solutions Manual Neil E. Schore, 2002-08-02 *Fundamentals of Chemistry, Study Guide* James E. Brady, John R. Holum, 1988-04-20 This Third Edition of the widely used fundamentals textbook for science majors maintains the conversational writing style that made the previous editions so popular while including up to date treatments of important and current topics Emphasizes descriptive chemistry chemical reactions and properties while maintaining a solid treatment of chemical principles Common chemicals are used whenever possible as examples in both theoretical discussions and in problems and exercises Incorporates many pedagogical aids each chapter begins with a brief table of contents and each section begins with a preview of topics covered Chapters include frequent margin comments figures and photographs

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