

EQUILIBRIUM IN CHEMICAL SYSTEMS

Name: _____

Use the following information to answer the **next two** questions.

The following chemical equation represents a reaction that occurs when soft drinks are carbonated.



- Which statement describes a carbonated beverage at equilibrium (in a soft drink bottle)?
 - The contents are in a closed system and have variable macroscopic properties.
 - The contents are in an open system and have variable macroscopic properties.
 - The contents are in a closed system and have constant macroscopic properties.
 - The contents are in an open system and have constant macroscopic properties.
- What occurs when this reaction is at equilibrium?
 - The decomposition of $\text{H}_2\text{CO}_{3(aq)}$ begins.
 - The formation of $\text{H}_2\text{CO}_{3(aq)}$ stops.
 - The decomposition of $\text{H}_2\text{CO}_{3(aq)}$ stops.
 - The rates of formation and decomposition of $\text{H}_2\text{CO}_{3(aq)}$ are equal.
- The quantities of reactants and products in chemical systems can be related as follows.
 - More reactants than products
 - More products than reactants
 - Equal reactants and products

Which relationship(s) can be true for chemical systems that are at equilibrium?

- III only
 - II only
 - I, II, and III
 - I and II
- A beaker contains a saturated aqueous solution of lead (II) chromate in contact with the solid compound. The system is represented by the chemical equation:



At 25°C , the solubility of lead (II) chromate is:

- moderate
 - very high
 - impossible to determine from the information given
 - very low
- Chemistry teachers often use the reaction between copper metal and aqueous silver nitrate to demonstrate a single replacement reaction. The net ionic equation for this reaction follows:

Chemistry Post Lab Answers Exploring Equilibrium

John A. Garraty, Mark C. Carnes



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