

Unit 2 Vocabulary - The Chemistry of Life

Nucleus	The positively charged core of hydrogen (H^+), found in acids.
Covalent Bond	Negatively charged molecule of one oxygen bonded to one hydrogen (OH^-), found in bases.
pH Scale	A molecule that exists in all three states of matter (solid, liquid and gas) on Earth, is the primary fluid of all living organisms.
Isotope	Tells us the number of atoms of each element in a compound, i.e. $NaCl$ is one sodium bonded to one chlorine and forming sodium chloride or salt.
ionic Bond	The transfer of valence electrons between atoms, formed between a metal and a nonmetal, the strongest (most polar) of all bond types.
Base	A process that involves the changing from one substance into something completely new or different.
Ion	A particle smaller than an atom, protons, neutrons and electrons.
Atoms	When electrons are bound to each other forming a compound (covalent or ionic) are the most common type.
Chemical Reaction	The amount of a substance in a specified space.
Acid	Any substance that forms hydroxide ions (OH^-) in water, has a pH greater than 7.
Subatomic Particle	The positively charged center of an atom, contains protons (p^+) and neutrons (neutrons).
Chemical Bond	A measure of how acidic or basic a solution is, ranges from 0 to 14, 7 is neutral, <7 is acidic, >7 is basic.
Concentration	Atoms of the same element that have a different number of neutrons in the nucleus.
Hydrogen Ion	The smallest particle of something, the basic building blocks of all matter.
Hydroxide Ion	Chemical bond formed when two atoms combine by the sharing of pairs of valence electrons.
Water	Any atom that has a charge, created by the loss (p^+) or gain (e^-) of electrons.
Chemical Formula	Any substance that forms hydrogen (H^+) ions in water, has a pH less than 7.

Chemistry Of Life Vocabulary Practice Answers

Tim Collins



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