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## #11 Qualitative Analysis Scheme of Main Group and Transition Metal Cations\*

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The following experiment is intended to continue the introduction of qualitative analysis through the identification of various main group and transition metal cations. Recall that the analysis of a substance can be extremely difficult given the fact that there are thousands of possibilities for reactions with some similar results. For example, you may have seen from the Group I qualitative analysis that  $\text{Ag}^+$ ,  $\text{Pb}^{+2}$ , and  $\text{Hg}_2^{+2}$  all produce a white precipitate when  $\text{Cl}^-$  is added.

In this experiment, the ions initially chosen were limited to aluminum, the alkaline earth metals, and certain first-row transition metals with the goal of minimizing the amount of hazardous waste generated. Analysis of both a known and unknown solution will test for  $\text{Al}^{+3}$ ,  $\text{Ba}^{+2}$ ,  $\text{Ca}^{+2}$ ,  $\text{Co}^{+2}$ ,  $\text{Cu}^{+2}$ ,  $\text{Fe}^{+3}$ ,  $\text{Mg}^{+2}$ ,  $\text{Mn}^{+2}$ , and  $\text{Ni}^{+2}$ . The ions are roughly divided into three groups. The group A ions have very insoluble oxides or hydroxides. The group B ions are poorly coordinated by  $\text{NH}_3$  and have very insoluble carbonates; magnesium is included in group B even though it is precipitated separately. Group C metal ions are strongly coordinated by  $\text{NH}_3$ . This prevents their precipitation in the previous steps.

Note that almost all of the solutions can safely go down the sink in moderate amounts. The barium sulfate should be collected, but its storage and disposal should not pose a serious problem in the small amounts used in the tests. The organic compounds used to test for the transition metal ions should be handled with care as they are irritants and suspected carcinogens. Once in solution, they are used in small amounts that should not pose a serious hazard.

### Procedure

#### Separation of Ion Groups: Unknown and Known (with each group)

1. Place 1 mL of your test solution in a medium test tube.
2. Add 12 drops of 6 M  $\text{HNO}_3$ , 3 mL of 6 M  $\text{NH}_3(\text{aq})$ , and stir. Centrifuge and decant. The solid may contain  $\text{Al}(\text{OH})_3$ ,  $\text{Fe}(\text{OH})_3$ , and  $\text{Mn}(\text{OH})_2$ . *Save the solid for Group A analysis.* Use the solution in the next step.
3. To the solution from step 2, add 2 mL of 1 M  $(\text{NH}_4)_2\text{CO}_3$  solution and stir. Centrifuge and decant. The solid may contain  $\text{BaCO}_3$  and  $\text{CaCO}_3$ . *Save the solid for Group B analysis.* Use the solution in the next step.

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\* Adapted from Peet, John T. A Short Qualitative Analysis Scheme without Hazardous Wastes

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