



## CALCIUM AND MAGNESIUM DETERMINATION

## OBJECTIVES

1. Apply the correct technique in the performance of calcium measurement in a given sample.
2. Analyze the laboratory results in terms of acceptability, accuracy, precision, and reliability.
3. Identify sources of error that may lead to false elevation or decrease in the calcium level.
4. Correlate laboratory results with clinical significance.

## CALCIUM DETERMINATION

## Materials

- Spectrophotometer and Cuvets
- Tissue
- Parafilm
- Serological pipets/automatic pipets
- Pipet tips
- Rubber aspirator
- Specimens (Control and unknown – serum or plasma)
- Test tube
- Centrifuge
- Calcium reagent kit

## Reagents

BUFFER	LYSINE BUFFER (pH 11.1)
Reagent	Sodium azide
	8-Hydroxyquinoline
	O-Cresolphthalein complexone
	Hydrochloric acid
Standard	Calcium: 8 mg/dL or 2 mmol/L
	Sodium azide

## Principle of the Method

- Calcium ions react with o-cresolphthalein complexone in an alkaline medium to form a purple colored complex. The absorbance of this complex is proportional to the calcium concentration in the sample.

## Assay Requirements

- Wavelength – 570 nm
- Optical path – 1 cm
- Temperature – 37°C
- Measurement – against water blank
- Preparation of Working Reagent:
  - Add color reagent to buffer in equal volumes as required and allow to stand for 10 minutes at room temperature before use.

## Procedure

## Pipetting Scheme: semi-micro

Pipet into cuvetts	B	S	C <sub>s</sub>	C <sub>p</sub>	U
Standard solution		0.02 mL			
Control normal serum			0.02 mL		
Control pathologic serum				0.02 mL	
Patient's serum					0.02 mL
Working reagent	1.0 mL	1.0 mL	1.0 mL	1.0 mL	1.0 mL

- Mix and measure the absorbance of sample ( $\Delta A_{sx}$ ) and standard ( $\Delta A_{STD}$ ) against the reagent blank within 5 to 50 minutes.

## CALCULATION

$$C = 8 \times \frac{A_{\text{sample}}}{A_{\text{standard}}} \quad \text{in mg/dL}$$

$$C = 2 \times \frac{A_{\text{sample}}}{A_{\text{standard}}} \quad \text{in mmol/L}$$

- **Reference range:** 2.15 – 2.50 mmol/L  
8.8– 10 mg/dL

# Chemistry Lab 2 Sem Guide

**University of Washington, Wash.)**  
**University of Washington (Seattle**

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web a how many milliliters of nitrogen can be made from 13 l of chlorine and 10.0 l of ammonia gas at stp  $1 \text{ l } 10.0 \text{ l } \text{NH}_3 \times 2 \text{ l } \text{NH}_3 \rightarrow 5.00 \text{ l } \text{N}_2$   $1 \text{ l } 13 \text{ l } \text{Cl}_2 \times 3 \text{ l } \text{Cl}_2 \rightarrow 4 \text{ l } \text{N}_2$  answer  $4.3 \times 10^3 \text{ ml } \text{N}_2$  b how many grams of chlorine must react to produce 16 l of nitrogen gas at 1.2 atm and 23°C  $\text{N}_2$   $1.2 \text{ atm} \times 16 \text{ l}$

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web he determined that if certain gases that are products and reactions in a chemical reaction are measured at the same conditions temperature and pressure then the volume of gas consumed produced is equal to the ratio between the gases or the ratio of the coefficients

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web this worksheet provides practice in stoichiometry problems involving gases at standard conditions stp and non standard conditions the relevant formulas for calculations are  $pV$  and  $n = \frac{V}{V_m}$  where  $V_m = 22.4 \text{ dm}^3 \text{ mol}^{-1}$  at stp  $t = 0^\circ\text{C}$   $p = 1 \text{ atm}$  the second half of the worksheet looks at problems relating to gas densities where the following

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web 1 carbon monoxide reacts with oxygen to produce carbon dioxide answer the following questions for the reaction of 1.0 L of carbon monoxide and oxygen at stp a write the balanced equation 20d or 200 b how many liters of  $\text{O}_2$  are required to react

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web may 28 2020 if gasoline contains 84.2% carbon by mass and has a density of 0.8205 g/mL determine the mass of carbon dioxide produced during a 500 mile trip 3.785 liters per gallon answer 1.28  $\times 10^5$  g  $\text{CO}_2$

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