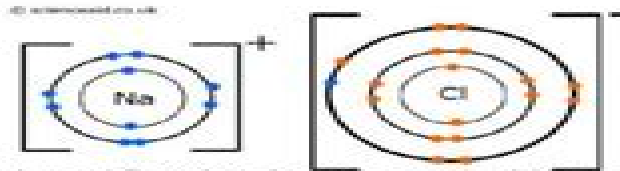


Chemical Bonding

Ionic Bonding

Ionic Compounds: Composed of anions and cations, these compounds do not form molecules. The chemical formulas for these are in the simplest ratio. Components are combined so that amount of positive and negative charge is equal. More likely than not forms a crystalline solid.



Think of it as an awkward first date between two middle schoolers: *no chemistry (so punny) no compound!*

*Ionic compounds form so that each atom, by **gaining or losing** electrons, has an octet of electrons in its highest occupied energy level.*

Formation: Metals lose electrons to form positively-charged cations, whereas nonmetals gain electrons to form negatively-charged anions.

The goal of the formation of a compound can be seen through electron configurations. Whatever adds up to make that of a noble gas (electron configuration) is desired to stabilize.

Ions come together due to electrostatic attraction, since opposites attract.

Formation (continued):

Attractive Forces: (What is attracted together) Oppositely charged ions, nuclei and electron clouds of adjacent ions.

Repulsive Forces: (What is **not** attracted together) Like-charged ions, electrons of adjacent ions.

The distance between the ions creates a balance between those forces; ions minimize their potential energy by combining in an arrangement called a crystal lattice. The specific lattice pattern created depends on the charges and

Chemistry Ionic Bonding Study Guide Answers

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