

Name: _____ Date: _____ Period: _____

pH Worksheet

1. What is pH a measure of?
2. What is the equation used for finding pH
3. What is the equation that relates to pH and pOH?
4. Complete the following table

$[H_3O^+]$	$[OH^-]$	pH	pOH	Acidic/Basic?
$1.0 \times 10^{-7} \text{ M}$	1×10^{-7}		5	
$2.4 \times 10^{-11} \text{ M}$		12.6	1.4	
	$5.62 \times 10^{-11} \text{ M}$	3.75	10.25	
$2.82 \times 10^{-9} \text{ M}$	$3.55 \times 10^{-6} \text{ M}$	8.55		

5. What would be the **pH** of each of the following:

- | | | | |
|------------------------------|-------|------------------------|-------|
| a) 0.0010 M HCl | _____ | g) 0.024 M HCl | _____ |
| b) 0.0010 M HNO ₃ | _____ | h) 0.075 M KOH | _____ |
| c) 0.010 M NaOH | _____ | i) 0.000034 M HCl | _____ |
| d) 0.0035 M HCl | _____ | j) 0.000000000001M HCl | _____ |
| e) 1.0 M HBr | _____ | | |
| f) 1.0 M KOH | _____ | | |

6. A 2.63 g NaOH are dissolved in 156 mL of solution. Determine the NaOH concentration & the pH.

7. List 3 strong acids and explain why these acids are considered strong acids.

8. List 3 weak acids and explain why these acids are considered weak acids.

9. List 2 strong bases and explain why these bases are considered strong bases.

10. List 1 weak base and explain why it is considered a weak base.

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ML Yell



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