

Gases Test SCH3U

Name:

 K °C °F °A Assessment of Learning (Product)

Multiple Choice. Circle the most correct answer. (15K)

1. Convert 120 K to °C.

- A) 393 °C B) -153 °C C) 880 °C D) 153 °C

2. Why do we assume that gas particles experience no intermolecular forces?

- A) Because they are moving too quickly B) Because they are small.
C) Because gas particles are a long distance from each other D) None of the above

3. The mass of a fixed quantity of a gas:

- A) increases when the volume increases, at constant temperature
B) decreases when the absolute temperature decreases, at constant temperature
C) does not change at any temperature and pressure
D) decreases as the density increases
E) decreases when the volume decreases, at constant pressure

4. Which rigid cylinder contains the same number of gas molecules at STP as a 2.0- liter rigid cylinder containing H_2 (g) at STP?

- A) 2.0- L cylinder of CH_4 (g)
B) 1.5- L cylinder of NH_3 (g)
C) 4.0- L cylinder of He (g)
D) 1.0- L cylinder of O_2 (g)

5. Which statement describes the particles of an ideal gas based on the kinetic molecular theory?

- A) The gas particles are relatively far apart and have negligible volume.
B) The gas particles are in constant, nonlinear motion.
C) The gas particles have attractive forces between them.
D) When gas particles have collisions they lose energy.



6. Which phase depicted above does not have a definite shape or volume?

- A) The one in the middle B) The one on the left
C) The one in the middle and the one on the right D) The one on the right

Chemistry Gases Unit Study Guide Answers

Jill Meryl Levy



Chemistry Gases Unit Study Guide Answers:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Hazmat Chemistry Study Guide (Second Edition) Jill Meryl Levy, 2011

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