

Recommended Questions for Unit 7 Gases (chapter 5): ALL book questions are recommended!!!! Here are a few samples:

1. What properties define a gas at a molecular level?
2. What effect will the following changes have on the volume of a fixed amount of gas?
 - a. An increase in pressure at constant temperature
 - b. An increase in pressure coupled with an increase in temperature
3. What effect will the following changes have on the pressure of a fixed amount of gas?
 - a. An increase in temperature at constant volume
 - b. An decrease in temperature coupled with a decrease in volume
4. Which of the following would not result in an increase in the volume of a fixed gas?
 - a. An increase in temperature at constant pressure
 - b. A decrease in temperature at constant pressure
5. Containers A and B are at the same pressure and temperature, but the volume of A is greater than that of B. Indicate whether a given gas would have the same or different densities in the two containers. If the densities are different, in which container is the density greater?
6. What is the pressure within a 10.0 L vessel containing 0.250 mol each of oxygen, hydrogen, and carbon dioxide at 298 K?
7. What is the mole fraction of oxygen in the mixture containing 0.25 mol of each oxygen, hydrogen, and carbon dioxide?
8. A comparison is made at STP of 0.50 mol $\text{H}_2(\text{g})$ and 1.0 mol $\text{He}(\text{g})$. Complete the statement: the two gases will a) have equal average molecular kinetic energies, b) have equal average molecular speeds, c) occupy equal volumes, d) have equal effusion rates.
9. A mixture of 5.0×10^{-3} mol $\text{H}_2(\text{g})$ and 5.0×10^{-3} mol $\text{SO}_2(\text{g})$ at 25°C is introduced into a 10.0-L container having a "pinhole" leak. After a period of time, the mole fraction of $\text{H}_2(\text{g})$ in the remaining gas mixture a) is the same as in the original mixture, b) is less than in the original mixture, c) is greater than the original mixture, d) cannot be determined relative to the original mixture
10. BE ABLE TO ANALYZE DIAGRAMS OF BAROMETERS, CLOSED-ENDED MANOMETERS, AND OPEN-ENDED MANOMETERS both qualitatively and quantitatively
11. Calculate the height of a column of water ($d = 1.00 \text{ g/mL}$) that exerts the same pressure as a column of mercury ($d = 13.6 \text{ g/mL}$) 760 mm high.
12. Without doing calculations, arrange the drawings in P5.5 so that the pressures denoted in red are in increasing order
13. A helium-filled party balloon has a volume of 4.50 L at sea level, where the atmospheric pressure is 748 Torr. Assuming the temperature remains constant, what will be the volume of the balloon when it is taken to a mountain resort at an altitude of 2500 m, where the atmospheric pressure is 557 Torr?
14. A gas is enclosed in a cylinder fitted with a piston. The volume of the gas is 2.00 L at 398 Torr. The piston is moved to increase the gas pressure to 5.15 atm. Which of the following is a reasonable value for the volume of the gas at the greater pressure? 0.20 L, 0.40 L, 1.00 L, or 16.0 L

Chemistry Gases Study Guide Answers

**Minu Gupta Bhowon, Sabina Jhaumeer-
Laulloo, Henri Li Kam Wah, Ponnadurai
Ramasami**

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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28

The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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