



CHEMISTRY

CHEAT SHEET



NAMING SIMPLE COMPOUNDS

NAMING SIMPLE COMPOUNDS CHEAT SHEET

Naming Compounds with Their Formulas
(Show the steps for naming the compound)



Steps for Naming:

1. Identify the cation (metal) and anion (non-metal).
2. Write the cation name first, followed by the anion name.
3. Use the appropriate suffix for the anion (usually -ide).
4. If the cation is a transition metal, include its oxidation state in Roman numerals in parentheses.
5. Check the formula to ensure the charges balance.
6. Write the final name.

Binary Compounds

Formula	Name
NaCl	Sodium Chloride
CaO	Calcium Oxide
Fe ₂ O ₃	Iron(III) Oxide
Al ₂ S ₃	Aluminum Sulfide
SnCl ₄	Tin(IV) Chloride
PbO ₂	Lead(IV) Oxide
Cr ₂ O ₃	Chromium(III) Oxide
Mn ₂ O ₇	Manganese(VII) Oxide
Co ₃ O ₂	Cobalt(II) Oxide
Ni ₃ O ₂	Nickel(II) Oxide
Fe ₃ O ₄	Iron(II,III) Oxide
SnO ₂	Tin(IV) Oxide

Common Ions and Their Names

Formula	Name
H ⁺	Hydrogen Ion
OH ⁻	Hydroxide Ion
Cl ⁻	Chloride Ion
Br ⁻	Bromide Ion
I ⁻	Iodide Ion
NO ₃ ⁻	Nitrate Ion
SO ₄ ²⁻	Sulfate Ion
PO ₄ ³⁻	Phosphate Ion
CO ₃ ²⁻	Carbonate Ion
SiO ₄ ⁴⁻	Silicate Ion

Naming Compounds with a Central and Terminal



Carbon Dioxide

Steps for Naming:

1. Identify the central atom and the terminal atom.
2. Write the central atom name first, followed by the terminal atom name.
3. Use the appropriate prefix for the number of terminal atoms.
4. Check the formula to ensure the charges balance.
5. Write the final name.

Prefixes:

Prefix	Number	Prefix	Number
mono	1	hepta	7
di	2	octa	8
tri	3	nona	9
tetra	4	deca	10
penta	5	undeca	11
hexa	6	dodeca	12

Common Names for Specific Compounds

Formula	Common Name	Formula	Common Name
NaCl	Salt	NaOH	Sodium Hydroxide
HCl	Hydrochloric Acid	Ca(OH) ₂	Slaked Lime
HNO ₃	Nitric Acid	CH ₃ COOH	Vinegar
H ₂ SO ₄	Sulfuric Acid	Na ₂ CO ₃	Soda Ash

Chemistry Foundation Naming Compounds Study Guide

Michael Seilmaier



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