

Chapter 1

Some Basic Concepts of Chemistry

Question 1.1:

Calculate the molar mass of the following:

- (i) H_2O (ii) CO_2 (iii) CH_4

Answer 1.1:

- (i) H_2O :

The molecular mass of water, H_2O
= $(2 \times \text{Atomic mass of hydrogen}) + (1 \times \text{Atomic mass of oxygen})$
= $[2(1.0084) + 1(16.00 \text{ u})]$
= $2.016 \text{ u} + 16.00 \text{ u} = 18.016 = 18.02 \text{ u}$

- (ii) CO_2 :

The molecular mass of carbon dioxide, CO_2
= $(1 \times \text{Atomic mass of carbon}) + (2 \times \text{Atomic mass of oxygen})$
= $[1(12.011 \text{ u}) + 2(16.00 \text{ u})]$
= $12.011 \text{ u} + 32.00 \text{ u} = 44.01 \text{ u}$

- (iii) CH_4 :

The molecular mass of methane, CH_4
= $(1 \times \text{Atomic mass of carbon}) + (4 \times \text{Atomic mass of hydrogen})$
= $[1(12.011 \text{ u}) + 4(1.008 \text{ u})]$
= $12.011 \text{ u} + 4.032 \text{ u} = 16.043 \text{ u}$

Question 1.2:

Calculate the mass per cent of different elements present in sodium sulphate (Na_2SO_4).

Answer 1.2:

The molecular formula of sodium sulphate is (Na_2SO_4).

Molar mass of $\text{Na}_2\text{SO}_4 = [(2 \times 23.0) + (32.066) + 4(16.00)] = 142.066 \text{ g}$

Mass percent of an element = $\frac{\text{Mass of that element in the compound}}{\text{Molar Mass of the compound}} \times 100$

\therefore Mass percent of sodium: $\frac{46.0 \text{ g}}{142.066 \text{ g}} \times 100 = 32.379 = 32.4\%$

Mass percent of sulphur: $\frac{32.066 \text{ g}}{142.066 \text{ g}} \times 100 = 22.57 = 22.6\%$

Mass percent of oxygen: $\frac{46.0 \text{ g}}{142.066 \text{ g}} \times 100 = 42.049 = 45.05\%$

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