Chemistry 338 2nd Semester Final Exam Review Sheet

District Assessment

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten	24 W	74	74	110	72	184	+2

- Name the compound.
 - a. Fe₂O₂ iron (III) oxide
- b. Cl₂O₈ dichlorine octaoxide

- Write the formula.
 - a. Trisulfur Dinitride SiNi
 - b. Zinc Hydroxide Zn(OH)₂
 - e. Tin (IV) oxide SnO2
- 3. The pressure of a gas is 843 mm Hg at 24°C. What will the pressure be at standard temperature? Volume is constant.

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$
 $\frac{843mmHg}{297K} = \frac{P_2}{273K}$

$$P_2 = 775 \text{ mm Hg}$$

4. The volume of a gas is \$1.7 mL when the pressure is 119 kPa and the temperature is 73°C. What is the volume at STP?

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$
 $\frac{(119kPa)(81.7mL)}{346K} = \frac{(101.3kPa)(V_2)}{273K}$ $V_2 = 75.7 \text{ mL}$

$$V_2 = 75.7 \text{ mL}$$

- Between Be and Ba, which has a larger ionization energy? Larger atomic radius? Larger ionization energy - Be-Larger atomic radius - Ba
- Balance the equation & classify the type of reaction:

single replacement

Predict the products & classify the type of reaction:

$$H_3PO_4 + 3 \text{ NaOH} \rightarrow 3 \text{ H}_3O + \text{ Na}_3PO_4$$

double replacement

Calculate the Molar Mass: (NH₄)₂CO₃

96 g/mole

Convert 2.75 moles of N2O3 to molecules

$$\frac{2.75moles}{1} \times \frac{6.02x10^{\circ} \cdot 23molecules}{1mole} = 1.66x10^{24} \text{ molecules N}_{2}O_{3}$$

Solutions

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M. J. Satriana, Montana State
University (Missoula), State University
of Montana (Missoula), University of
Montana--Missoula

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