



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In chemistry, it refers to a reaction where the forward and reverse rates are equal, and the concentrations of reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the net change is zero.</p>	<p>Equilibrium Constants</p> <p>K_c (Concentration): $K_c = \frac{[Products]}{[Reactants]}$</p> <p>$K_p$ (Pressure): $K_p = \frac{P_{Products}}{P_{Reactants}}$</p> <p>$K_a$ (Acid Dissociation): $K_a = \frac{[H^+][A^-]}{[HA]}$</p> <p>$K_b$ (Base Dissociation): $K_b = \frac{[OH^-][BH^+]}{[B]}$</p>													
<p>Equilibrium Constant, K_c</p> <p>For a general reaction: $aA + bB \rightleftharpoons cC + dD$</p> <p>$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$</p> <p>Units: K_c is dimensionless.</p>	<p>Equilibrium Constant, K_p</p> <table><tr><th>Reaction</th><th>Equilibrium Constant</th></tr><tr><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$</td><td>$K_p = \frac{P_{NH_3}^2}{P_{N_2} P_{H_2}^3}$</td></tr><tr><td>$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$</td><td>$K_p = \frac{P_{SO_3}^2}{P_{SO_2}^2 P_{O_2}}$</td></tr><tr><td>$H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$</td><td>$K_p = \frac{P_{HI}^2}{P_{H_2} P_{I_2}}$</td></tr></table>		Reaction	Equilibrium Constant	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$	$K_p = \frac{P_{NH_3}^2}{P_{N_2} P_{H_2}^3}$	$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$	$K_p = \frac{P_{SO_3}^2}{P_{SO_2}^2 P_{O_2}}$	$H_2(g) + I_2(g) \rightleftharpoons 2HI(g)$	$K_p = \frac{P_{HI}^2}{P_{H_2} P_{I_2}}$				
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<p>Equilibrium Constant, K_a</p> <p>For a weak acid: $HA \rightleftharpoons H^+ + A^-$</p> <p>$K_a = \frac{[H^+][A^-]}{[HA]}$</p> <p>Units: K_a is dimensionless.</p>	<p>Equilibrium Constant, K_b</p> <table><tr><th>Reaction</th><th>Equilibrium Constant</th></tr><tr><td>$B + H_2O \rightleftharpoons BH^+ + OH^-$</td><td>$K_b = \frac{[BH^+][OH^-]}{[B]}$</td></tr><tr><td>$CO_3^{2-} + H_2O \rightleftharpoons HCO_3^- + OH^-$</td><td>$K_b = \frac{[HCO_3^-][OH^-]}{[CO_3^{2-}]}$</td></tr><tr><td>$CH_3COO^- + H_2O \rightleftharpoons CH_3COOH + OH^-$</td><td>$K_b = \frac{[CH_3COOH][OH^-]}{[CH_3COO^-]}$</td></tr></table>		Reaction	Equilibrium Constant	$B + H_2O \rightleftharpoons BH^+ + OH^-$	$K_b = \frac{[BH^+][OH^-]}{[B]}$	$CO_3^{2-} + H_2O \rightleftharpoons HCO_3^- + OH^-$	$K_b = \frac{[HCO_3^-][OH^-]}{[CO_3^{2-}]}$	$CH_3COO^- + H_2O \rightleftharpoons CH_3COOH + OH^-$	$K_b = \frac{[CH_3COOH][OH^-]}{[CH_3COO^-]}$				
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<p>Le Chatelier's Principle: The effect of changing conditions on equilibrium</p> <table><tr><th>Condition</th><th>Effect on Equilibrium</th><th>Example</th></tr><tr><td>Concentration</td><td><ul style="list-style-type: none">Increasing concentration of a reactant shifts the equilibrium to the right.Increasing concentration of a product shifts the equilibrium to the left.</td><td>$Fe^{3+} + SCN^- \rightleftharpoons FeSCN^{2+}$ Adding more Fe^{3+} or SCN^- shifts the equilibrium to the right, forming more $FeSCN^{2+}$.</td></tr><tr><td>Pressure</td><td><ul style="list-style-type: none">Increasing pressure shifts the equilibrium to the side with fewer moles of gas.Decreasing pressure shifts the equilibrium to the side with more moles of gas.</td><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ Increasing pressure shifts the equilibrium to the right, forming more NH_3.</td></tr><tr><td>Temperature</td><td><ul style="list-style-type: none">Increasing temperature shifts the equilibrium to the endothermic side.Decreasing temperature shifts the equilibrium to the exothermic side.</td><td>$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$ This reaction is exothermic. Increasing temperature shifts the equilibrium to the left, forming more SO_2 and O_2.</td></tr></table>			Condition	Effect on Equilibrium	Example	Concentration	<ul style="list-style-type: none">Increasing concentration of a reactant shifts the equilibrium to the right.Increasing concentration of a product shifts the equilibrium to the left.	$Fe^{3+} + SCN^- \rightleftharpoons FeSCN^{2+}$ Adding more Fe^{3+} or SCN^- shifts the equilibrium to the right, forming more $FeSCN^{2+}$.	Pressure	<ul style="list-style-type: none">Increasing pressure shifts the equilibrium to the side with fewer moles of gas.Decreasing pressure shifts the equilibrium to the side with more moles of gas.	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ Increasing pressure shifts the equilibrium to the right, forming more NH_3 .	Temperature	<ul style="list-style-type: none">Increasing temperature shifts the equilibrium to the endothermic side.Decreasing temperature shifts the equilibrium to the exothermic side.	$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$ This reaction is exothermic. Increasing temperature shifts the equilibrium to the left, forming more SO_2 and O_2 .
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Chemistry Equilibrium Study Guide

Martha Mackin



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clearly state the purpose of your request in the opening paragraph 03 provide necessary details to support your request such

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