

Chapter 10 Outline

Chemical Quantities

- **Section 10.1 – The Mole: A Measurement of Matter**

- You often measure the amount of something by _____, by _____, or by _____.
- A _____ (mol) of a substance is _____ representative particles of that substance.
- 6.02×10^{23} is called _____ number.
- 1 mole = _____
- A _____ refers to the species present in a substance: usually _____.
- Elements normally exist as _____, but 7 elements exist as _____ molecules: _____.

- **Sample Problem**

- How many moles is 2.80×10^{24} atoms of silicon?

- **Practice Problems**

- How many moles is 2.17×10^{23} representative particles of bromine?
- How many molecules are in 2.12 mol of propane? (m/c = molecules)

- **Sample Problem**

- How many atoms are in 1.14 mol SO_3 ?

- **Practice Problems**

- How many moles are in 4.65×10^{24} molecules of NO_2 ?

- How many atoms are in 4.33 mol magnesium sulfate?

- The _____ of an element expressed in grams is the mass of a _____ of the element.
- The mass of a _____ of an element is the _____.
- To calculate the _____ of a compound, find the number of _____ of each _____ in one mole of the compound. Then _____ the masses of the elements in the compound.

- **Sample Problem**

- What is the molar mass of PCl_3 ?

- **Practice Problems**

- What is the molar mass of sodium hydrogen carbonate?
- What is the mass of calcium nitrate?

Chemistry Chemical Quantities Study Guide

L Cohen



Chemistry Chemical Quantities Study Guide:

Study Guide for Introductory Chemistry Iris Stovall, 1990 *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02

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