

CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS

Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.

1 mol = a molar mass of an atom/molecule (g/mol)

EXAMPLE: What is the mass of 5.00 moles of water H_2O ?

$$\begin{aligned} \text{STEP 1: } 2 \text{ H} &= 2 \text{ x } (1.01) = 2.02 \\ \text{O} &= 1 \text{ x } (16.00) = 16.00 \end{aligned}$$

$$H_2O = 18.02 \text{ g}$$

$$\text{STEP 2: } 5.00 \text{ moles } H_2O \times \frac{18.02 \text{ g } H_2O}{1 \text{ mole } H_2O} = 90.10 \text{ g } H_2O$$

NOW YOU TRY ONE: What is the mass of 0.50 moles of calcium carbonate $CaCO_3$?

$$\begin{aligned} \text{STEP 1: } Ca &= \\ C &= \\ 3O &= \end{aligned}$$

$$CaCO_3 =$$

$$\text{STEP 2: } 7 \text{ g } CaCO_3 =$$

USE A SEPARATE SHEET OF PAPER TO SOLVE THE FOLLOWING PROBLEMS. SHOW YOUR WORK. ROUND MOLAR MASSES TO TWO PLACES AFTER THE DECIMAL. ADD UNITS.

- How many grams are there in 5.00 moles of lead Pb ?
310.8 g Pb
- How many grams are there in 2.00 moles of sulfuric acid H_2SO_4 ?
196.16 g H_2SO_4
- How many grams are there in 0.250 moles of sodium hydroxide $NaOH$?
10 g NaOH
- How many grams are there in 2.50 moles of potassium nitrate KNO_3 ?
252.75 g KNO_3
- How many grams are there in 10.0 moles of lithium carbonate Li_2CO_3 ?
738.9 g Li_2CO_3

Now that you know how to find the mass of one mole of a substance you can easily find the number of moles there are in a given mass of the substance.

EXAMPLE: How many moles of calcium chloride are there in 333 grams of calcium chloride ($CaCl_2$)?

$$\begin{aligned} \text{STEP 1: } Ca &= 1 \text{ x } (40.08) = 40.08 \\ 2 \text{ Cl} &= 2 \text{ x } (35.45) = 70.90 \\ CaCl_2 &= 110.98 \text{ g} \end{aligned}$$

$$\text{STEP 2: } 333 \text{ grams } CaCl_2 \times \frac{1 \text{ mole } CaCl_2}{110.98 \text{ g } CaCl_2} = 3.00 \text{ mole } CaCl_2$$

USE THE SAME PAPER AS THE ABOVE PROBLEMS TO SOLVE THE FOLLOWING. SHOW YOUR WORK AND PUT UNITS ON EACH ANSWER!

- How many moles of silver nitrate are there in 80.00 grams of silver nitrate $AgNO_3$?
0.47 mol $AgNO_3$
- How many moles of phosphoric acid are there in 658 grams of phosphoric acid H_3PO_4 ?
6.71 mol H_3PO_4
- How many moles of tin (II) fluoride are there in 908 grams of tin (II) fluoride SnF_2 ?
5.79 mol SnF_2
- How many moles of hydrogen peroxide are there in 1000.0 grams of hydrogen peroxide H_2O_2 ?
29.39 mol H_2O_2
- How many moles of magnesium chloride are there in 148 grams of magnesium chloride $MgCl_2$?
1.55 mol $MgCl_2$

#1-5: You are given the number of moles and asked to find the grams (mass)

MOLS \rightarrow GRAMS

#6-10: You are given the grams (mass) and asked to find the number of mols

GRAMS \rightarrow MOLS

Chemistry B Moles Packet

American Chemical Society

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the American Chemical Society text.

Chemistry B Moles Packet:

Chemistry John Olmsted, Gregory M. Williams, 1997 Textbook outlining concepts of molecular science Chemistry of Cement, 1962 Excel HSC Chemistry Jim Stamell, 2011 ISBN 9781741252996 AUTHOR Jim Stamell RRP 39 95 PAGES 428 pp SPECIFICATION Softcover perfect bound 280 mm x 210 mm STATUS New edition PUBLICATION DATE April 2008 The EXCEL HSC Chemistry guide is directly linked to the syllabus with every single dot point of the HSC Chemistry syllabus appearing in the margin of the book You can write in the guide so your study is focused and your notes are structured This guide comes in a brand new format that makes even better use of your study time up to date coverage of the core topics plus 3 Option topics Industrial Chemistry Shipwrecks Corrosion and Conservation and Forensic Chemistry this guide is organised just like the HSC syllabus so the students learn to section the theoretical part is under routine headings and the students section the practical part is under headings like First hand Second hand and Investigations and Problem Solving %this way you will be able to see at a glance what the theoretical and practical work is all main headings in each chapter 1 1 2 1 etc are directly from the syllabus word for word %this way you can easily match the Excel guide to the syllabus an alphabetical list of all the key definitions and concepts you should know from each chapter %an efficient way of learning all the definitions in one go chapter syllabus checklist with every single dot point listed in checklist form for each chapter %a fantastic way of testing that you know all the work hundreds of key concept questions with answers %questions that test your recall of knowledge in each chapter HSC type questions for every section in each chapter with clock icons to tell you how much time you will have to answer the questions in the HSC %this way you can test yourself on HSC type questions under HSC type time pressure an examiner maximiser feature ticks to show the mark distribution and answers to all HSC type questions %all you need to answer HSC type questions two sample HSC papers with an examiner maximiser feature plus answers %not one but two up to date sample papers the Excel syllabus summary notes a detachable section at the end of the guide where every single dot point of each chapter is summarised for you% a comprehensive and compact summary of the whole course in 32 pages *Chemistry* John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant **Comprehensive Chemistry XI** Dr. B. Kapila, S. K. Khanna, 2010-11 Comprehensive chemistry according to the new syllabus prescribed by Central Board of Secondary Education CBSE Chemistry in the Community (ChemCom) American Chemical Society, 2011-06-17 Touted as the most successful NSF funded project published

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Rickard, 2010-12-28 CHEMISTRY **Foundations of College Chemistry** Morris Hein, Susan Arena, 2013-01-01 Learning the fundamentals of chemistry can be a difficult task to undertake for health professionals For over 35 years Foundations of College Chemistry Alternate 14th Edition has helped readers master the chemistry skills they need to succeed It provides them with clear and logical explanations of chemical concepts and problem solving They ll learn how to apply concepts with the help of worked out examples In addition Chemistry in Action features and conceptual questions checks brings together the understanding of chemistry and relates chemistry to things health professionals experience on a regular basis

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