#### Types of Chemical Reactions Worksheet #1

#### HJ/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 2Mg + O<sub>2</sub> → 2 MgO Synthesis
- Cal<sub>2</sub> + Cl<sub>2</sub> → CaCl<sub>2</sub> + I<sub>2</sub> Single Replacement
- 3) 3KOH + AlCl<sub>3</sub> → Al(OH)<sub>3</sub> + 3KCl Double Replacement
- C + O<sub>2</sub> → CO<sub>2</sub> Synthesis
- \*5) Ca(OH)<sub>2</sub> + 2HCl → CaCl<sub>2</sub> + 2H<sub>2</sub>O Double Replacement
- 6) 2KClO<sub>3</sub> → 2KCl+3 O<sub>2</sub> Decomposition
- 3CuSO<sub>4</sub> + 2Al → Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> + 3Cu Single Replacement
- 8) Na<sub>2</sub>S + 2AgNO<sub>3</sub> → 2NaNO<sub>3</sub> + Ag<sub>2</sub>S Double Replacement
- 9) 3H<sub>2</sub> + N<sub>2</sub> → 2NH<sub>3</sub> Synthesis
- \*10) 2Na + 2H2O -> 2NaOH + H2 Single Replacement

\*Remember: "H2O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- Double replacement: Na<sub>2</sub>CrO<sub>4</sub> + PbCl<sub>2</sub> → 2 NaCl + PbCrO<sub>4</sub>
- Single replacement: Cl<sub>2</sub> + 2 NaBr → 2 NaCl + Br<sub>2</sub>
- Decomposition: Mg(ClO<sub>3</sub>)<sub>2</sub> → MgCl<sub>2</sub> + 3 O<sub>2</sub>
- 14) Synthesis:  $2 H_2 + O_2 \rightarrow 2 H_2O$
- 15) Double replacement: 3 Ca(OH)<sub>2</sub> + 2 FeCl<sub>3</sub> → 2 Fe(OH)<sub>3</sub> + 3 CaCl<sub>2</sub>
- 16) Single replacement:  $Fe + Cu(NO_3)_2 \rightarrow Fe(NO_3)_2 + Cu$  [Assume  $Fe^{2+}$ ]
- 17) Decomposition: 2 Hg<sub>2</sub>O → 4 Hg + O<sub>2</sub>
- 18) Synthesis:  $S + O_2 \rightarrow SO_2$
- Double replacement: AgNO<sub>3</sub> + KI → AgI + KNO<sub>3</sub>
- 20) Single replacement: Cu + 2 AgNO<sub>3</sub> → 2 Ag + Cu(NO<sub>3</sub>)<sub>2</sub> [Copper (II) is used here]

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