

pH AND pOH

Chapter 14 Section 14.2

The pH of a solution indicates how acidic or basic that solution is.

pH ranges of 0-7 = acidic

7 = neutral

7-14 = basic

Since $pH + pOH = 14$ for all aqueous solutions, the pOH of a solution can be calculated using either value.

$pH = 4 \Rightarrow pOH =$

So if $pH = 10$ then $pOH =$

$pOH = 4 \Rightarrow pH =$

So if $pOH = 10$ then $pH =$

Together $pH + pOH =$

Complete the following chart.

	pH	pH	pOH	pOH	Acids or Bases
1.	normal	0	normal	14	Acidic
2.	10^{-10} M	1	10^{-1} M	13	acidic
3.	10^{-9} M	2	normal	12	acidic
4.	normal	3	10^{-12} M	11	acidic
5.	10^{-12} M	4	10^{-11} M	10	acidic
6.	10^{-11} M	5	10^{-10} M	9	acidic
7.	10^{-9} M	6	normal	8	acidic
8.	normal	7	10^{-7} M	7	neutral
9.	10^{-7} M	8	10^{-6} M	6	basic
10.	10^{-6} M	9	10^{-5} M	5	basic

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Richard Bailey



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