

Chem 12 Practice Solubility Test

1. Which combination of factors will affect the rate of the following reaction?



- A. Temperature and surface area only
- B. Temperature and concentration only
- C. Concentration and surface area only
- D. Temperature, concentration, and surface area

2. Consider the following reaction:



The rate of decomposition of the oxalate ion is increased by

- A. Adding NaOH
- B. Removing CO_2
- C. Adding a catalyst
- D. Decreasing the pressure

3. An equilibrium system shifts left when the temperature is increased. The forward rate is

- A. Exothermic and ΔH is positive.
- B. Exothermic and ΔH is negative.
- C. Endothermic and ΔH is positive.
- D. Endothermic and ΔH is negative.

4. The value of the K_{eq} can be changed by

- A. adding a catalyst.
- B. changing the temperature.
- C. changing the reactant concentration.
- D. changing the volume of the container

5. Consider the following equilibrium: $2\text{NOCl}_{(g)} \rightleftharpoons 2\text{NO}_{(g)} + \text{Cl}_{2(g)}$

In a 1.0 L container at equilibrium there are 1.0 mol NOCl, 0.70 mol NO and 0.40 mol Cl_2 . At constant temperature and volume, 0.10 mol NOCl is added. The concentrations in the "new" equilibrium in comparison to the concentrations in the "old" equilibrium are

	[NOCl]	[NO]	[Cl ₂]
A	new = old	new = old	new = old
B	new > old	new > old	new > old
C	new < old	new < old	new > old
D	new < old	new > old	new > old

Chemistry 12 Solubility Questions

Jeffrey Gaffney, Nancy Marley



Chemistry 12 Solubility Questions:

This book delves into Chemistry 12 Solubility Questions. Chemistry 12 Solubility Questions is a vital topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemistry 12 Solubility Questions, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Chemistry 12 Solubility Questions
- Chapter 2: Essential Elements of Chemistry 12 Solubility Questions
- Chapter 3: Chemistry 12 Solubility Questions in Everyday Life
- Chapter 4: Chemistry 12 Solubility Questions in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Chemistry 12 Solubility Questions. The first chapter will explore what Chemistry 12 Solubility Questions is, why Chemistry 12 Solubility Questions is vital, and how to effectively learn about Chemistry 12 Solubility Questions.

3. In chapter 2, the author will delve into the foundational concepts of Chemistry 12 Solubility Questions. This chapter will elucidate the essential principles that need to be understood to grasp Chemistry 12 Solubility Questions in its entirety.

4. In chapter 3, the author will examine the practical applications of Chemistry 12 Solubility Questions in daily life. This chapter will showcase real-world examples of how Chemistry 12 Solubility Questions can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of Chemistry 12 Solubility Questions in specific contexts. This chapter will explore how Chemistry 12 Solubility Questions is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Chemistry 12 Solubility Questions. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Chemistry 12 Solubility Questions.

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