## MCQ FOR BOARD EXAMINATION XI Chapter - 01

## STOICHIOMETRY

- If the volume occupied by oxygen gas (O<sub>2</sub>) at STP is 44.8dm<sup>3</sup>, the number of molecules of O<sub>2</sub> in the Vessels are:
- (a)  $12.01 \times 10^{24}$ (b)  $12.01 \times 10^{25}$ (c)  $12.01 \times 10^{23}$ (d)  $12.01 \times 10^{22}$  The number of Carbon atoms in 1 mole of sugar (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>) are approximately: (a)  $12.01 \times 10^{23}$ (b)  $12.01 \times 10^{24}$
- (c)  $12.01 \times 10^{23}$ (d)  $72 \times 10^{23}$  In the reaction 2Na + 2H<sub>2</sub>O → 2NaOH + H<sub>2</sub>, if 23g of Na reacts with excess of water, the volume of hydrogen gas (H2) liberated at STP should be: (b) 2.44 dm<sup>3</sup> (a) 11.2 dm<sup>3</sup> (c) 10.60 dm<sup>3</sup> (d) 44.8 dm2
- 4. Which of the following sample of substance contains the same number of atoms as that of 20g calcium:
  - (a) 16g S (b) 32g O (c) 39g K (d) 48g Mg
- 5. Which of the following statement is incorrect?
  - (a) The mass of 1 mole Cl2 gas is 35.5g
  - (b) One mole of H<sub>2</sub> gas contains 6.01 × 10<sup>23</sup> molecules of H<sub>2</sub>
  - (c) Number of atoms in 23g Na and 24g of Mg are equal.
  - (d) One moles of O<sub>2</sub> at S.T.P occupy 22.4 dm<sup>3</sup>
- For Avogadro's number, this statement is incorrect:
  - (a) It is the number of particles in one mole of any substance
  - (b) Its numerical value is  $6.01 \times 10^{23}$
  - (c) Its value change if temperature increases (d) Its value change if number of moles increases
- 7. The minimum number of moles are present in:
- - (a) I dm of methode gas at STP (b) 5 dm³ of helium gas at STP (c) 10 dm, of hydrogen gas at STP (d) 22.4 dm3 of chlorine gas at STP
- 8. Number of atoms in 60g of Carbon are:
- (c)  $6.01 \times 10^{25}$ (d)  $6.01 \times 10^{24}$ (a) 12:01 x 10<sup>23</sup> (b)  $3.01 \times 10^{24}$
- Maximum number of molecules present in the following sample of gas:
  - (a) 100g O2 (c) 100g CO<sub>2</sub>

(b) 100g of CH<sub>4</sub> (d) 100g of Cl<sub>2</sub>

(d)  $6.0 \times 10^{3}$ 

(d) 6.02 x 10<sup>-6</sup>

(d)  $3.3 \times 10^{-3}$ 

- 10. Generally actual yield is:
  - (a) Greater than theoretical yield
  - (b) Less than theoretical yield (c) Equal to the theoretical yield
- (d) Some times greater and sometimes less than theoretical yield 11. The short hand disruption of a very large or small number in the power of 10 is called:
- (a) Exponential notation (b) Rounding off (c) Significant figure (d) All of them
- 12. 60000 is equal to:
  - (a)  $6 \times 105^{\circ}$ (b) 6.00 x 10° (c) 60 x 10°
- 13. 602000 in exponential notation may be expressed as:
- (a)  $6.02 \times 10^{\circ}$ (b) 6.02 x 10° (c) 0.602 x 10° 14. 330000 in exponential notation may be expressed as:
- (a)  $3.3 \times 10^{\circ}$ (b)  $0.33 \times 10^{5}$ (c)  $3.3 \times 10^{-5}$
- 15. The number 683 x 10<sup>-7</sup> expressed in scientific notation as (c)  $6.83 \times 10^{-11}$ (d) 6.83 x 10<sup>-12</sup> (a) 68.3 x 10° (b) 6.83 x 10<sup>-5</sup>
- 16. The reactant which does not completely consume in the reaction and some of its amount remains unused at the end of reaction.
  - (a) Reactant in excess (b) Limiting reactant (c) Product in excess (d) Reaction
- I mole of Gold (Au) 197g how many no of atom.
  - (a)  $6.02 \times 10^{23}$  atom (b)  $12.02 \times 10^{23}$  atom (c)  $6.02 \times 10^{25}$  atom (d)  $12.02 \times 10^{25}$  atom

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