

Chemistry Semester 2 Review

Unit 9: The Mole

Examples to Practice: (Show all work for calculations)

1. What are the following conversion factors called? Include the value of any that are constants.
 A.) moles & particles
 B.) grams & moles

C.) moles & liters (assume STP)

2. Order the following quantities from smallest to largest.

A. 0.5 mol
 B. pair
 C. 5

D. 6.02×10^{23}
 E. dozen
 F. four moles

smallest _____ → largest _____

3. Determine the conversion factors for the following problems.

a. How many moles of calcium are in 825 g calcium (Ca)?

b. How many atoms are in 2.15 moles of water?

c. How many atoms of copper are in 12.0 grams?

4. Show the relationship between empirical and molecular formulas by placing the given formulas in the table below.

Empirical Formula	Compound	Molecular Formula
	acetylene	
	benzene	
	ethylene	
	butane	
	cyclohexane	
	formaldehyde	
	acetic acid	
	glyceraldehyde	

Given formulas			
C_2H_2	$C_6H_6O_3$	CH_2O	C_6H_6
$C_2H_4O_2$	C_2H_2	CH_2	C_6H_6
C_2H_6	OH	CH_2O	

5. What is the molar mass of $K_2Cr_2O_7$?

10. Determine the number of molecules in 16.75 g of H_2O ?

6. What is the molar mass of $Fe_2(CO_3)_3$?

(HONORS ONLY)

7. What is the molar mass of $CoCl_2 \cdot 6H_2O$?

11. Find the formula for a hydrate with 48.8% $MgSO_4$ & 51.2% H_2O

8. How many moles of NH_3 contain 1.75×10^{24} molecules of?

12. What is the % by mass of C in $Pb(C_2H_3O_2)_2$?

9. How many molecules are in 0.26 mol of CO_2 ?

13. Find the empirical and molecular formula for a compound that contains 49.98g carbon and 10.47g hydrogen? (58.12 g/mol)

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