

# Organic Chemistry

## Questions and Answers

2018  
Edition

ene is large. Aromaticity in naphthalene is also large. Aromaticity has come to mean a stabilization resulting from (p-orbital) overlap in a pi-type system. As the examples are large or small.

Let's consider  $C_{60}$ :

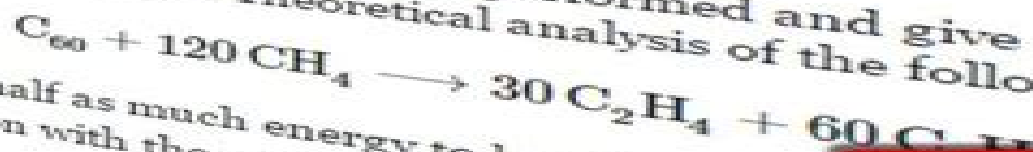
Bond alternation is often taken as a sign of non-aromatic systems. In  $C_{60}$ , bond lengths, ~1.4 and 1.45 angstroms. However, this variation is on the order of that found in polycyclic aromatic hydrocarbons, and less than that observed in non-aromatic systems. Conclusion: aromatic, but less so than benzene.

Magnetic properties are related to electron delocalization and are often used to measure aromaticity. Both experiment and calculations suggest the existence of ring currents (diamagnetic and paramagnetic) in  $C_{60}$ .

Conclusion: Although analysis is complex, analysis is consistent with  $C_{60}$  being aromatic.

Reactivity - Substitution reactions are not possible as no hydrogens are attached. If an anion or radical is added to  $C_{60}$  the electron(s) are not delocalized over the entire structure. However, most addition reactions are reversible suggesting that  $C_{60}$  is not as aromatic as benzene.

Resonance energy calculations have been performed and give conflicting results. Most suggest a small stabilization. Theoretical analysis of the following isomerization reaction suggests that it only took half as much energy to break a C-C bond as in benzene. Conclusion: Some aromaticity is present.



by George Duckett

# Chemistry Questions Answers George Duckett

**Lei Shi**



## **Chemistry Questions Answers George Duckett:**

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