



# CHEMISTRY

## CHEAT SHEET



# NAMING SIMPLE COMPOUNDS

## NAMING SIMPLE COMPOUNDS CHEAT SHEET

Naming Compounds with Their Formulas  
(Show the steps for naming the compound)



**Steps for Naming:**

1. Identify the cation (metal) and anion (non-metal).
2. Name the cation first, followed by the anion.
3. Use the appropriate suffix for the anion (usually -ide).
4. If the cation is a transition metal, include its oxidation state in parentheses.
5. Check the formula to ensure the charges balance.
6. Write the final name.

### Binary Compounds

Formula	Name
NaCl	Sodium Chloride
CaO	Calcium Oxide
Fe <sub>2</sub> O <sub>3</sub>	Iron(III) Oxide
Al <sub>2</sub> S <sub>3</sub>	Aluminum Sulfide
SnCl <sub>4</sub>	Tin(IV) Chloride
PbO <sub>2</sub>	Lead(IV) Oxide
K <sub>2</sub> SO <sub>4</sub>	Potassium Sulfate
Li <sub>3</sub> N	Lithium Nitride
Ag <sub>2</sub> Br	Silver Bromide
Cr <sub>2</sub> O <sub>3</sub>	Chromium(III) Oxide

### Common Ions and Their Names

Formula	Name
H <sup>+</sup>	Hydrogen Ion
OH <sup>-</sup>	Hydroxide Ion
Cl <sup>-</sup>	Chloride Ion
Br <sup>-</sup>	Bromide Ion
I <sup>-</sup>	Iodide Ion
NO <sub>3</sub> <sup>-</sup>	Nitrate Ion
SO <sub>4</sub> <sup>2-</sup>	Sulfate Ion
PO <sub>4</sub> <sup>3-</sup>	Phosphate Ion
CO <sub>3</sub> <sup>2-</sup>	Carbonate Ion
CH <sub>3</sub> COO <sup>-</sup>	Acetate Ion

Naming Compounds with a Central and Monatomic



Carbon Dioxide

**Steps for Naming:**

1. Identify the central atom and the monatomic anion.
2. Name the central atom first, followed by the anion.
3. Use the appropriate suffix for the anion (usually -ide).
4. If the central atom is a transition metal, include its oxidation state in parentheses.
5. Check the formula to ensure the charges balance.
6. Write the final name.

**Naming Practice**

Formula	Name	Formula	Name
NaCl	Sodium Chloride	CaO	Calcium Oxide
Fe <sub>2</sub> O <sub>3</sub>	Iron(III) Oxide	Al <sub>2</sub> S <sub>3</sub>	Aluminum Sulfide
SnCl <sub>4</sub>	Tin(IV) Chloride	PbO <sub>2</sub>	Lead(IV) Oxide
K <sub>2</sub> SO <sub>4</sub>	Potassium Sulfate	Li <sub>3</sub> N	Lithium Nitride
Ag <sub>2</sub> Br	Silver Bromide	Cr <sub>2</sub> O <sub>3</sub>	Chromium(III) Oxide

### Common Names for Specific Compounds

Formula	Common Name	Formula	Common Name
NaCl	Salt	NaOH	Sodium Hydroxide
H <sub>2</sub> O	Water	CO <sub>2</sub>	Carbon Dioxide
H <sub>2</sub> SO <sub>4</sub>	Sulfuric Acid	CH <sub>3</sub> COOH	Acetic Acid
Na <sub>2</sub> CO <sub>3</sub>	Soda Ash	Na <sub>2</sub> SO <sub>4</sub>	Sodium Sulfate

# Chemistry Naming Compounds Study Guide

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## Chemistry Naming Compounds Study Guide:

*Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

**Study Guide and Solutions Manual** Neil E. Schore, 2002-08-02 *Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

**Organic Chemistry Study Guide with Solutions Manual** Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the

errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

**Barron's Science 360: A Complete Study Guide to Chemistry with Online Practice** Barron's Educational Series, Mark Kernion, Joseph A. Mascetta, 2021-09-07 provides a complete guide to the fundamentals of chemistry Page 4 of cover

*Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl* Paul B. Kelter, 1993 *Organic Chemistry, Student Study Guide and Solutions Manual* David R. Klein, 2017-01-04 This is the Student Study Guide and Solutions Manual to accompany Organic Chemistry 3e Organic Chemistry 3rd Edition is not merely a compilation of principles but rather it is a disciplined method of thought and analysis Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems Readers must learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry Existing textbooks provide extensive coverage of the principles but there is far less emphasis on the skills needed to actually solve problems

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