

## Ionic Compounds and Metals

### Section 7.1 Ion Formation

pages 206–209

#### Section 7.1 Assessment

page 209

1. **Compare** the stability of a lithium atom with that of its ion,  $\text{Li}^+$ .

The  $\text{Li}^+$  ion is more stable because it has a complete octet.

2. **Describe** two different causes of the force of attraction in a chemical bond.

The two causes are the attraction between the positive nucleus of one atom and the negative electrons of another atom, and the attraction between positive ions and negative ions.

3. **Apply** Why are all of the elements from group 18 relatively unreactive, whereas those in group 17 are very reactive?

The group 18 elements, known as noble gases, have complete outer energy levels and do not easily form ions. The group 17 elements are highly reactive because they need only to gain a single electron to form an octet.

4. **Summarize** ionic bond formation by correctly pairing these terms: *cation*, *anion*, *electron gain*, *electron loss*.

(anion, electron gain), (cation, electron loss)

5. **Apply** Write out the electron configuration for each atom. Then predict the change that must occur in each to achieve a noble gas configuration.

- a. nitrogen

$[\text{He}]2s^22p^3$ ; gain 3 electrons ( $3^-$  ion) or lose 5 electrons ( $5^+$  ion)

- b. sulfur

$[\text{Ne}]3s^23p^4$ ; gain 2 electrons ( $2^-$  ion)

- c. barium

$[\text{Xe}]6s^2$ ; lose 2 electrons ( $2^+$  ion)

- d. lithium

$[\text{He}]2s^1$ ; lose 1 electron ( $1^+$  ion)

6. **Model** Draw models to represent the formation of the positive calcium ion and the negative bromide ion.

Models should show that the calcium atom loses two electrons, forming  $\text{Ca}^{2+}$ , and that bromine gains one electron, forming  $\text{Br}^-$ . The models should also show the addition of energy to form the calcium ion, and the release of energy to form the bromide ion.

### Section 7.2 Ionic Bonds and Ionic Compounds

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#### Practice Problems

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Explain how an ionic compound forms from these elements.

7. sodium and nitrogen

Three Na atoms each lose  $1 e^-$ , forming  $1^+$  ions. One N atom gains  $3 e^-$ , forming a  $3^-$  ion. The ions attract, forming  $\text{Na}_3\text{N}$ .

$$3 \text{ Na ions } \left( \frac{1+}{\text{Na ion}} \right) + 1 \text{ N ion } \left( \frac{3-}{\text{N ion}} \right) \\ = 3(1+) + 1(3-) = 0$$

The overall charge on one formula unit of  $\text{Na}_3\text{N}$  is zero.

8. lithium and oxygen

Two Li atoms each lose  $1 e^-$ , forming  $1^+$  ions. One O atom gains  $2 e^-$ , forming a  $2^-$  ion. The ions attract, forming  $\text{Li}_2\text{O}$ .

$$2 \text{ Li ions } \left( \frac{1+}{\text{Li ion}} \right) + 1 \text{ O ion } \left( \frac{2-}{\text{O ion}} \right) \\ = 2(1+) + 1(2-) = 0$$

The overall charge on one formula unit of  $\text{Li}_2\text{O}$  is zero.

# Chemistry Matter And Change Chapter 7 Study Guide Answers

**Léo Pomar**



## **Chemistry Matter And Change Chapter 7 Study Guide Answers:**

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *Chemistry: Molecules, Matter, and Change Media Activities Book* Loretta Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfield-Tacher, 2000-01-15 The Media Activity Book MAB for Jones Atkins Chemistry molecules matter and change contains chapters with lists and descriptions of some of the media available as you study the chapter Each activity begin with a specific textbook reference Then you are given a time estimate of how long it will take to use the media An M media icon in the margin of the textbook means that media exists to support that area of text The media is found in three different places on the website and on two CDs **O Level Chemistry Questions and Answers PDF** Arshad Iqbal, The O Level Chemistry Quiz Questions and Answers PDF IGCSE GCSE Chemistry Competitive Exam Questions Chapter 1 14 Practice Tests Class 9 10 Chemistry Textbook Questions for Beginners includes revision guide for problem solving with hundreds of solved questions O Level Chemistry Questions and Answers PDF book covers basic concepts analytical and practical assessment tests O Level Chemistry Quiz PDF book helps to practice test questions from exam prep notes The O Level Chemistry Quiz Questions and Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved tests O Level Chemistry Questions and Answers PDF Free download chapter 1 a book covers solved common questions and answers on chapters Acids and bases chemical bonding and structure chemical formulae and equations electricity electricity and chemicals elements compounds mixtures energy from chemicals experimental chemistry methods of purification particles of matter redox reactions salts and identification of ions and gases speed of reaction and structure of atom tests for school and college revision guide Chemistry Interview Questions and Answers PDF Download free eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The IGCSE GCSE Chemistry Interview Questions Chapter 1 14 PDF book includes high school question papers to review practice tests for exams O Level Chemistry Practice Tests a textbook s revision guide with chapters tests for IGCSE NEET MCAT GRE GMAT SAT ACT competitive exam O Level Chemistry Questions Bank Chapter 1 14 PDF book covers problem solving exam tests from chemistry textbook and practical eBook chapter wise as Chapter 1 Acids and Bases Questions Chapter 2 Chemical Bonding and Structure Questions Chapter 3 Chemical Formulae and Equations Questions Chapter 4 Electricity Questions

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scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using evaluation criteria developed for the guide The criteria reflect and incorporate goals and principles of the National Science Education Standards The annotations designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter on educational software and multimedia programs chapters on books about science and teaching directories and guides to science trade books and periodicals for teachers and students Another section features institutional resources One chapter lists about 600 science centers museums and zoos where teachers can take middle school students for interactive science experiences Another chapter describes nearly 140 professional associations and U S government agencies that offer resources and assistance Authoritative extensive and thoroughly indexed and the only guide of its kind Resources for Teaching Middle School Science will be the most used book on the shelf for science teachers school administrators teacher trainers science curriculum specialists advocates of hands on science teaching and concerned parents

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