# Chemistry 100 Final Study Guide Key

#### Chapter 1: Introduction

- a. Physical vs. Chemical properties
  - Physical properties: shiny, metallic, malleable, ductile, conductive. Chemical property: reactive with water.
- b. Units of measurement
  - 1) Meters, liters, grams (with any prefixes) are metric system. All others are U.S.
  - 21 5.68 x 10-7 46.300.000
- c. Density
  - 1) 2,47g/mL

#### Chapter 3: Atomic Structure

- a. Historical Perspective
  - List the major scientific accomplishment of the following scientists:
    - (i) Experiments in electrochemistry described ions and the electrical nature of atoms
    - (ii) Cathode-ray tube experiments led to the proposal of the "Plum pudding" model of atoms
    - (iii) Oil-drop experiments calculated the mass of electrons
    - (iv) Gold foil experiments proved atoms had a dense, positively-charged nucleus and empty space
    - (v) Proposed an atomic model with a nucleus and electrons circling like planets
    - (vi) Proposed that electrons exist in a "cloud" or probability region around the nucleus
- b. Isotopes
  - Isotopes of an element have the same number of protons, but different numbers of neutrons. Regular hydrogen has 1 proton and no neutrons. Isotopes could have 1 proton and 1 neutron (deuterium) or 1 proton and 2 neutrons (tritium).
- Structure of the atom
  - Protons and neutrons are in the nucleus. Electrons are in regions of space surrounding the nucleus.
  - 3 energy levels, 4 sublevels, 6 orbitals, 11 electrons
  - 3) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>2</sup>
- Electron configurations and trends on the Periodic Table
  - See the periodic Table.

#### . Chapter 5: Chemical bonds

- a. Electron-dot structures
  - Draw the electron-dot structure of an oxygen atom.
- . 0

- Octet Rule and lons.
  - Mg<sup>2+</sup>, Cir. N<sup>3-</sup>
- c. Ionic compounds
  - 1) Aluminum bromide, calcium sulfide, sodium nitride
- d. Covalent compounds
  - Sulfur dioxide, phosphorus trichloride, carbon tetrahydride
- e. Polyatomic molecules
  - Nitrate (NO<sub>2</sub>), chlorate (ClO<sub>2</sub>), carbonate (CO<sub>2</sub>), sulfate (SO<sub>2</sub>), chromate (CrO<sub>2</sub>), phosphate (PO<sub>2</sub>)
- Electron-dot formulas for molecules.
  - Answer will be given in class.
- vSEPR Model and Molecular shape

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