

HONORS STUDY GUIDE: MOLES TEST

Vocabulary:

Moles- 6.02×10^{23} of anything is equal to 1 mole. A counting unit.

Representative Particles- can stand for atoms, moles, formula units, etc.

Avagadro's number- 6.02×10^{23}

Percent composition- the percent by mass of each element in a compound.

Empirical Formula- the simplest form of an equation. The lowest whole number ratio.

Molecular Formula- the actual number of atoms in a formula. Not reduced.

Hydrate- ionic compound with water molecules trapped in it.

Anhydrous- Compound that does not contain water molecules.

Mole Conversions:

1. How many moles are in 28.6g of CH_2 ?

2.04 moles

2. What is the mass of 3.2 moles of H_2O ?

57.66 g

3. How many formula units are in 2.8×10^{23} moles of $\text{C}_6\text{H}_{12}\text{O}_6$?

1.68×10^{45} F.U.

4. How many moles are in 6.8×10^{23} atoms of Carbon?

1.13×10^{-44} moles

5. How many particles are in 24.3 g of Sodium chloride?

2.50×10^{23} particles

6. What is the mass of 5.8×10^{16} molecules of $\text{Mg}_3(\text{PO}_4)_2$?

2.53×10^{-5} g

Chemistry Honors Mole Conversion Study Guide

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