

Chemistry Finals Study Guide

Chapter 1 & 2 - Review of Basic Concepts

- **Describe the 3 states of matter.** The three states of matter are *solid*, *liquid*, and *gas*.
 - **Solid:** When atoms in a material are moving slow that they stick together and barely move, creating a hard object classified as a solid
 - **Liquid:** When atoms are moving slower than a gas state but faster than the solid state so they flow and are able to take the shape of their container. They have more energy than solid state atoms.
 - **Gas:** When atoms are moving faster than the other two states of matter (solid & liquid) that they are no longer pulled towards each other and roam freely without boundaries if not kept in a container.
- **Describe the difference between elements and compounds, mixtures.**
 - **Elements:** materials that are only made out of the same *type* of atoms and cannot be changed into simpler substances by a chemical change (ex: O_2)
 - **Compounds:** materials that are a mix of two or more element in a fixed proportion
 - **Mixtures:** a mixture of 2 or more pure substances.
 - **Heterogeneous Mixture:** can often be separated by filtration
 - **Homogenous Mixture:** can be separated by distillation, crystallization, and chromatography
- **Convert between Celsius and Kelvin.**
 - **Celsius to Kelvin:** $C = (^{\circ}K - 273.15)$
 - **Kelvin to Celsius:** $K = (^{\circ}C + 273.15)$
- **Round measured and calculated values to the correct number of significant digits.**
 - Significant Digit Rules: [\[CHEM\] Final 1 - Steps & Processes](#)
- **Use dimensional analysis to solve multi step equations, showing how units cancel**
[\[CHEM\] Final 1 - Steps & Processes](#)
- **What is the difference between precision and accuracy?**
 - **Precision:** Acquiring the same result again and again
 - **Accuracy:** A result that is close to the accepted value/actual answer

Chemistry Finals 2015 Study Guide Answers

Linda Anne Silvestri, Angela E. Silvestri

A red circular graphic with a gradient, appearing as a partial circle or a stylized 'C' shape, located to the right of the authors' names.

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Chemistry John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

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and under naturalistic conditions Interestingly the teams conducting these studies are already exemplars of partnerships between researchers and practitioners who are uniquely positioned as in between straddling the two worlds As a result these publications represent both the rigours of research and the pragmatism of reflective practice In forthcoming editions we will add to this collection a third type of publication design profiles These will present practitioner developed pedagogical designs at varying levels of abstraction to be held to scrutiny amongst practitioners instructional designers and researchers alike We hope by bringing these types of studies together in an open access format that we may contribute to the development of new forms of practitioner researcher interactions that promote co design in pedagogical innovation

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









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