

1. What is the proper symbol for an atom with 45 protons, 101 neutrons and 45 electrons?
A) $^{146}_{45}\text{Rh}$
B) $^{146}_{101}\text{Md}$
C) $^{146}_{45}\text{Rh}$
D) $^{146}_{101}\text{Fm}$
E) I am not sure.
2. What is the name of the molecular compound IF_7 ?
A) Moniodide fluoride
B) Moniodine heptafluoride
C) Iodine heptafluoride
D) Iodine fluoride
E) I am not sure.
3. Which element has the following electron configuration: $1s^2 2s^2 2p^4 3s^2 3p^4$?
A) N
B) S
C) P
D) Se
E) I am not sure.
4. Write the formula unit formed by copper (III) oxide.
A) CuO
B) Cu_2O
C) Cu_2O_3
D) Cu_3O_2
E) I am not sure.
5. Balance the following reaction:
 $__\text{C}_2\text{H}_6(\text{g}) + __\text{O}_2(\text{g}) \rightarrow __\text{CO}_2(\text{g}) + __\text{H}_2\text{O}(\text{g})$
A) 1, 3, 2, 1
B) 2, 5, 2, 2
C) 1, 3, 2, 2
D) 2, 6, 4, 4
E) I am not sure.
6. Element X has three natural isotopes: X-4 (4.240 amu), X-5 (5.025 amu), and X-7 (7.349 amu). Calculate the mass of element x to two decimal places given the abundance X-4 is 20.6 % and X-5 is 36.7 %.
A) 5.86 amu
B) 5.61 amu
C) 5.75 amu
D) 5.65 amu
E) I am not sure.
7. What is the charge for the cobalt ion in CoO ?
A) $1+$
B) $2+$
C) $2-$

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