

CHEMISTRY

EQUATIONS AND ANSWERS

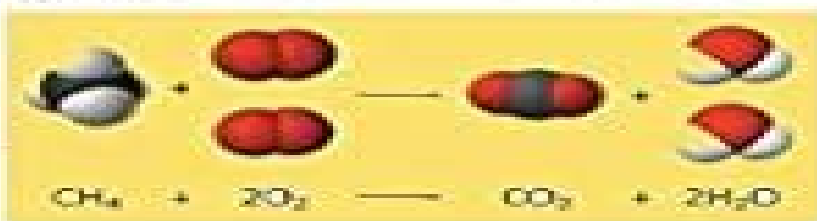
BALANCING CHEMICAL EQUATIONS

The law of conservation of mass dictates that the quantity of each element does not change in a chemical reaction. Thus, each side of the chemical equation must represent the same quantity of any particular element. Likewise, the charge is conserved in a chemical reaction. Therefore, the same charge must be present on both sides of the balanced equation.

One balances a chemical equation by changing the scalar number for each chemical formula. Simple chemical equations can be balanced by inspection, that is, by trial and error. Another technique involves solving a system of linear equations.

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One balances a chemical equation by changing the scalar number for each chemical formula. Simple chemical equations can be balanced by inspection, that is, by trial and error. Another technique involves solving a system of linear equations.



As seen from the equation $\text{CH}_4 + 2\text{O}_2 \longrightarrow \text{CO}_2 + 2\text{H}_2\text{O}$, a coefficient of 2 must be placed before the oxygen gas on the reactants side and before the water on the products side in order for, as per the law of conservation of mass, the quantity of each element does not change during the reaction.

As an example, seen in the above image, the burning of methane would be balanced by putting a coefficient of 1 before the CH_4 .



Since there is one carbon on each side of the arrow, the first atom (carbon) is balanced.

Looking at the next atom (hydrogen), the right-hand side has two atoms, while the left-hand side has four. To balance the hydrogens, 2 goes in front of the H_2O , which yields:



Inspection of the last atom to be balanced (oxygen) shows that the right-hand side has four atoms, while the left-hand side has two. It can be balanced by putting a 2 before O_2 , giving the balanced equation:



This equation does not have any coefficients in front of CH_4 and CO_2 , since a coefficient of 1 is dropped.

IONIC EQUATIONS

An ionic equation is a chemical equation in which electrolytes are written as dissociated ions. Ionic equations are used for single and double displacement reactions that occur in aqueous solutions. For example in the following precipitation reaction:



the full ionic equation is:



In this reaction, the Ca^{2+} and the NO_3^{-} ions remain in solution and are not part of the reaction. That is, these ions are identical on both the reactant and product side of the chemical equation. Because such ions do not participate in the reaction, they are called spectator ions. A net ionic equation is the full ionic equation from which the spectator ions have been removed. The net ionic equation of the preceding reaction is:



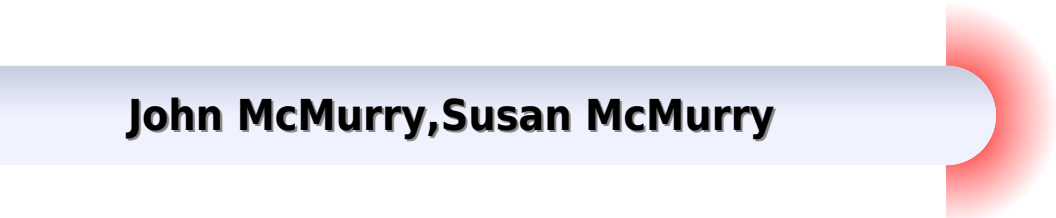
or, in reduced balanced form,



In a neutralization or acid/base reaction, the net ionic equation will usually be:

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and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Table of Contents Chemistry Equations And Answers Speedy Study Guides

1. Understanding the eBook Chemistry Equations And Answers Speedy Study Guides
 - The Rise of Digital Reading Chemistry Equations And Answers Speedy Study Guides
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Equations And Answers Speedy Study Guides
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Equations And Answers Speedy Study Guides
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Equations And Answers Speedy Study Guides
 - Personalized Recommendations
 - Chemistry Equations And Answers Speedy Study Guides User Reviews and Ratings
 - Chemistry Equations And Answers Speedy Study Guides and Bestseller Lists
5. Accessing Chemistry Equations And Answers Speedy Study Guides Free and Paid eBooks

- Chemistry Equations And Answers Speedy Study Guides Public Domain eBooks
 - Chemistry Equations And Answers Speedy Study Guides eBook Subscription Services
 - Chemistry Equations And Answers Speedy Study Guides Budget-Friendly Options
6. Navigating Chemistry Equations And Answers Speedy Study Guides eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Equations And Answers Speedy Study Guides Compatibility with Devices
 - Chemistry Equations And Answers Speedy Study Guides Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Equations And Answers Speedy Study Guides
 - Highlighting and Note-Taking Chemistry Equations And Answers Speedy Study Guides
 - Interactive Elements Chemistry Equations And Answers Speedy Study Guides
 8. Staying Engaged with Chemistry Equations And Answers Speedy Study Guides
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Equations And Answers Speedy Study Guides
 9. Balancing eBooks and Physical Books Chemistry Equations And Answers Speedy Study Guides
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Equations And Answers Speedy Study Guides
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Equations And Answers Speedy Study Guides
 - Setting Reading Goals Chemistry Equations And Answers Speedy Study Guides
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Equations And Answers Speedy Study Guides
 - Fact-Checking eBook Content of Chemistry Equations And Answers Speedy Study Guides
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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steering rangka komponen mesin sepeda motor serta fungsinya cylinder head crankcase cylinder block piston dan ring piston bearing kruk as crankshaft valve valve spring dan rocker arm flywheel bantalan aksial dan kedudukan katup poros bubungan dan pen piston push rod dan valve lifter bagian bagian motor serta fungsinya

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