

Reviewing Content

- 4.2. The solvent is the substance in which the solute is dissolved.
- 4.3. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
- 4.4. **solubility:** the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.
saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution:** a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible:** describes liquids that dissolve in each other. **immiscible:** describes liquids that are insoluble in each other.
- 4.5. Particles of solute crystallize.
- 4.6. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
- 4.7. 5.55×10^2 g AgNO₃.
- 4.8. Solubility increases with pressure.
- 4.9. a. 1.6×10^{-2} g/L
b. 4.7×10^{-2} g/L
- 5.0. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
- 5.1. Molarity is the number of moles of solute dissolved in one liter of solution.
a. 1.2M KCl
b. 3.3×10^{-1} M MgCl₂
- 5.2. 2.00×10^3 mL.
- 5.3. a. 5.0×10^{-2} mol NaCl, 29 g NaCl
b. 1.0 mol KNO₃, 1.0×10^{-2} g KNO₃
c. 2.5×10^2 mol CaCl₂, 2.8 g CaCl₂
- 5.4. a. 2.3×10^1 g NaCl
b. 2.0 g MgCl₂
- 5.5. a. 16% (v/v) ethanol
b. 63.6% (v/v) isopropyl alcohol

56. Colligative properties are properties of a solution that depend only on the number of solute particles: boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water
b. 1.5M KNO₃
c. 0.100M MgCl₂
58. The effective molality of the Ca(NO₃)₂ solution is 3m. The effective molality of the NaNO₃ solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g H₂O to 32.0 g CH₃OH.
63. a. 100.26°C
b. 101.54°C
64. a. -4.46°C
b. -2.2°C
65. a. -1.1°C
b. -0.74°C
c. -1.5°C

Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

Chemistry Chapter 16 Solutions Test B

Padhraic Smyth

Chemistry Chapter 16 Solutions Test B:

The book delves into Chemistry Chapter 16 Solutions Test B. Chemistry Chapter 16 Solutions Test B is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemistry Chapter 16 Solutions Test B, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Chemistry Chapter 16 Solutions Test B
- Chapter 2: Essential Elements of Chemistry Chapter 16 Solutions Test B
- Chapter 3: Chemistry Chapter 16 Solutions Test B in Everyday Life
- Chapter 4: Chemistry Chapter 16 Solutions Test B in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Chemistry Chapter 16 Solutions Test B. The first chapter will explore what Chemistry Chapter 16 Solutions Test B is, why Chemistry Chapter 16 Solutions Test B is vital, and how to effectively learn about Chemistry Chapter 16 Solutions Test B.
3. In chapter 2, the author will delve into the foundational concepts of Chemistry Chapter 16 Solutions Test B. The second chapter will elucidate the essential principles that need to be understood to grasp Chemistry Chapter 16 Solutions Test B in its entirety.
4. In chapter 3, the author will examine the practical applications of Chemistry Chapter 16 Solutions Test B in daily life. The third chapter will showcase real-world examples of how Chemistry Chapter 16 Solutions Test B can be effectively utilized in everyday scenarios.
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6. In chapter 5, this book will draw a conclusion about Chemistry Chapter 16 Solutions Test B. This chapter will summarize the key points that have been discussed throughout the book.

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