General Chemistry II Chapter 15 Study Guide

- S' pu > S' nond > S' mad
- For similar molecules, S' tends to increase with:
 - Increasing molar mass
 - Increased number of atoms in the formula

Standard entropy of a reaction (ΔS_{rea}) :

 ΔS_{res}: is the entropy change that occurs when all reactants and products are in their standard states:

$\Delta S'_{ren} = \Sigma m S'_{products} - \Sigma n S'_{reactants}$

where m and n are the mois of products and reactants, given by the coefficients in the balanced equation.

$$\sum mS^*_{products} = c \Delta S^*(c) + d \Delta S^*(0)$$

Remember to multiply AH of each substance by its coefficient.

Predicting the sign of ΔS_{con} :

- We can often predict the sign of ΔS_{rea} for processes that involve a change in the number of moles of gas
 - ΔS_{con} is positive if the amount of gas increases
 - ΔS_{ma} is negative if the amount of gas decreases
 - ΔS_{roe} is likely to be positive if a new structure forms that has more freedom of motion

Gibbs free energy:

- Gibbs free energy (G): combines the enthalpy and entropy of a system:
 - G = H TS
- Gibbs free energy change (AC): is a measure of the spontaneity of a process and of the useful energy available from it:
 - $\Delta G_{ex} = \Delta H_{ex} T\Delta S_{ex}$
- ΔG ≤ 0 for a spontaneous reaction (-ΔG)
- ΔG ≥ 0 for a nonspontaneous reaction (+ΔG)
- $\Delta G = 0$ for a process at equilibrium ($\Delta G = 0$)

Chemistry Chapter 15 Study Guide

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