

18) What are the two products that are **always** formed in a combustion reaction?

_____ + _____

19) When balancing a chemical reaction, what are the **large numbers** that we adjust? _____

20) When balancing a chemical reaction, what are the **small numbers** in molecules? _____

21) What does the **Law of Conservation of Mass** tell us about both sides of a reaction? Explain.

In the questions below, put a **checkmark** next to the equations that are **not** balanced correctly. ————



Use the following reaction to answer the questions below:



30) In the reaction above, what does the **(aq)** stand for next to H_2O_2 ? _____

31) In the reaction above, what does the **(l)** stand for next to H_2O ? _____

32) In the reaction above, what does the **(g)** stand for next to O_2 ? _____

33) Knowing this information above, what would an **(s)** symbol stand for? _____

34) In the reaction above, list the **reactant(s)**: _____, the **product(s)**: _____

35) What is the arrow called? _____

36) In the expression **3 NO₂**, which number is the **coefficient** and which number is the **subscript**?

3 = _____ 2 = _____

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Neil E. Schore



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