

Acids and Bases

Overview of Acids and Bases

Indicators: Acids and Bases

Acids: Donates its proton to base (proton donor)

Base: Donates its lone pair to acid (LBP)



Remember: (1) only H atoms, (2) only OH is a base

What about organic indicators and conjugate indicators?

Conjugate Acids and Bases

Acids are proton donors, Bases are proton acceptors.



General acid-base reactions of acid-base reactions:

to find the strong base, the base gets negative charge of the base and attacks the most acidic hydrogen of the acid.

to find the weak base, the base gets negative charge of the base and attacks the most acidic hydrogen of the acid.

Remember: -OH is a weak base, -OH⁻ is a strong base

What makes the Acid Strong? The strength of the acid is determined by the stability of the conjugate base. The more stable the conjugate base, the stronger the acid.

Acid Strength

Acids with a more stable conjugate base are stronger acids.



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Acid Strength and pKa

The Ka value of an acid (Ka) and its conjugate base (Kb) are related by:



The acid is stronger and the base is weaker.

To compare the acid, the pKa is used. The lower the pKa, the stronger the acid.

pKa = -log Ka

$$pK_a + pK_b = 14$$

Acids with a lower pKa are stronger acids.

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The Position of Equilibrium for an Acid-Base Reaction

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Choosing a Base to Deprotonate a Given Compound

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Chemistry Acid Base Study Guide

Leo J. Malone



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