

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
 (a) formula unit. (c) crystal.
 (b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
 (a) molecules. (c) dipoles.
 (b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
 (a) higher in potential energy. (c) equal in potential energy.
 (b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
 (a) Compound A is not an ionic compound.
 (b) It will be more difficult to break the bonds in compound A than those in compound B.
 (c) Compound B has larger crystals than compound A.
 (d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
 (a) stronger than the attractive forces among formula units in ionic bonding.
 (b) weaker than the attractive forces among formula units in ionic bonding.
 (c) approximately equal to the attractive forces among formula units in ionic bonding.
 (d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

- What type of energy best represents the strength of an ionic bond?

Lattice energy

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

Chemistry A Modern Course Answers

Jianjun Gao



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Table of Contents Chemistry A Modern Course Answers

1. Understanding the eBook Chemistry A Modern Course Answers
 - The Rise of Digital Reading Chemistry A Modern Course Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry A Modern Course Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry A Modern Course Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry A Modern Course Answers
 - Personalized Recommendations
 - Chemistry A Modern Course Answers User Reviews and Ratings
 - Chemistry A Modern Course Answers and Bestseller Lists
5. Accessing Chemistry A Modern Course Answers Free and Paid eBooks
 - Chemistry A Modern Course Answers Public Domain eBooks
 - Chemistry A Modern Course Answers eBook Subscription Services

- Chemistry A Modern Course Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Chemistry A Modern Course Answers Compatibility with Devices
 - Chemistry A Modern Course Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry A Modern Course Answers
 - Highlighting and Note-Taking Chemistry A Modern Course Answers
 - Interactive Elements Chemistry A Modern Course Answers
- 8. Staying Engaged with Chemistry A Modern Course Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry A Modern Course Answers
- 9. Balancing eBooks and Physical Books Chemistry A Modern Course Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry A Modern Course Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry A Modern Course Answers
 - Setting Reading Goals Chemistry A Modern Course Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry A Modern Course Answers
 - Fact-Checking eBook Content of Chemistry A Modern Course Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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