## **Experiment 17-Chemical Reactions Lab**

Since the Middle Ages, when ancient physicians attempted to find a magical substance that would cure all diseases, humans have been fascinated with chemical reactions. In order to effectively describe these changes, a system for and series of short hand notation was developed to aid ancient and now modern chemists classify chemical reactions.

The names of the basic building blocks of matter (elements) have historically interesting names such as Ytterbium, named after a cute little town in Sweden, and Plumbus, named after a metal (lead) that was used for water pipes in ancient Rome. The names of compounds, or combinations of different elements, are much more complicated. To simplify matters, chemists use one or two letter abbreviations for the elements. Writing a compound is simplified by using a chemical formula, or a series of symbols that represent the elements present in the compound.

Chemical compounds, as we will learn, have definite, whole number ratios of various. Typically, the atoms in a compound are listed in the formula in the order (from left to right) that the elements appear on the periodic table. The number ratios of the atoms in a compound are shown by using subscripts (little numbers that are lower than the rest of the text) just to the right of the element. Examples of formulas are shown below.

NaCl sodium chloride, with 1 sodium atom and 1 chlorine atom CO<sub>2</sub> carbon dioxide, with 1 carbon atom for every 2 oxygen atoms HCl hydrochloric acid, with 1 hydrogen atom for each chlorine atom

A chemical reaction involves elements and/or compounds combining, separating or changing partners in various ways. Again, describing such a complicated process in words is difficult, so a chemical reaction equation is used. The equation shows the chemical formula(s) of the starting compound(s), called reactants, and the final compound(s), called products. The reactants are always shown on the left, and the products on the right, and they are separated by an arrow to ensure that no one is confused on which direction the reaction is going. A general chemical reaction equation for a reaction called combination is shown below, and a specific example follows.

reactant A + reactant B  $\rightarrow$  product C reaction 1 C (s) + O<sub>2</sub> (g)  $\rightarrow$  CO<sub>2</sub> (g) reaction 2

If you would like to read the equation aloud, this is what the general reaction #1 would say: A reacts with B to make C. Equation 2 would read: Carbon reacts with oxygen to make carbon dioxide. The funny letters in parentheses in equation 2 represent the physical state that the matter is in, such as gas (g), liquid (l), solid (s), or dissolved in water (aq).

In order to aid chemists in recognizing and describing reactions, several types of reactions are classified into groups. Equation 2 is an example of a *combination* reaction, where two reactants combine to form 1 product. Examples of other types of reactions are listed below:

Decomposition, where one reactant decomposes or separates into 2 or more simpler products:

reactant A → product B + product C

 $H_2CO_3 \rightarrow CO_2(g) + H_2O(l)$ 

Carbonic acid decomposes in water to make carbon dioxide gas and water.

# **Chemical Reaction Lab Report Conclusion**

**RJ Alexander** 

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Interactions of Matter Christine Caputo, 2010 A look at how different elements interact in chemical reactions to form compounds with new properties CHEMISTRY EXPERIMENTS James Signorelli, 2014-09-19 Gifted and talented students and any student interested in pursuing a science major in college needs a rigorous program to prepare them while they are still in high school This book utilizes a format where the application of several disciplines science math and language arts principles are mandated Each lab concludes with either an essay or a detailed analysis of what happened and why it happened This format is based on the expectations of joining a university program or becoming an industrial science professional The ideal student lab report would be written in a lab research notebook and then the essay or final analysis is done on a word processor to allow for repeat editing and corrections The research notebook has all graph pages a title section and a place for the students and their assistants to sign and witness that exercise The basic mechanics of the lab report title purpose procedure diagrams data table math and calculations observations and graphs are handwritten into the book The conclusion is done on a word processor MS Word which allows the instructor to guide the student in writing and editing a complete essay using the MLA format When the final copy is completed the essay is printed and inserted into the lab notebook for grading At the end of the term the student has all their labs in one place for future reference These lab notebooks can be obtained for as little as 3 00 per book This is money well spent In our district the Board of Education buys the books for each student The BOE sees these books as expendable but necessary materials for all science and engineering instruction Everything You Need to Ace Chemistry in One Big Fat Notebook Workman Publishing, Jennifer Swanson, 2020-09-01 Chemistry No problem This Big Fat Notebook covers everything you need to know during a year of high school chemistry class breaking down one big bad subject into accessible units Learn to study better and get better grades using mnemonic devices definitions diagrams educational doodles and guizzes to recap it all Including Atoms elements compounds and mixtures The periodic table Quantum theory Bonding The mole Chemical reactions and calculations Gas laws Solubility pH scale Titrations Le Chatelier s principle and much more Forensics in Chemistry Sara McCubbins, Angela Codron, 2012 Forensics seems to have the unique ability to maintain student interest and promote content learning I still have students approach me from past years and ask about the forensics case and specific characters from the story I have never had a student come back to me and comment on that unit with the multiple choice test at the end from the Introduction to Forensics in Chemistry The Murder of Kirsten K How did Kirsten K s body wind up at the bottom of a lake and what do wedding cake ingredients soil samples radioactive decay bone age blood stains bullet matching and drug lab evidence reveal about whodunit These mysteries are at the core of this teacher resource book which meets the unique needs of high school chemistry classes in a highly memorable way The book makes forensic evidence the foundation of a series of eight hands on week long labs As you weave the labs throughout the year and students solve the case the narrative provides vivid lessons in

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