# CHEMICAL EQUILIBRIUM

# # Basic theory of equilibrium

- · Equilibrium represents the state ed a process in which the measurable proposties like-temp., pressure, colour, conc. of the system donot Show any change with the passage ob Hime.
- · equilibrium is a dyanamic process (1) conc. of reactant e product are chemical equilibrium can be oppround brow both sides.
- . The state ob equilibrium is not attacted by the presence ob catalast It only helps to attain the equilibrium state in secs or more Arme.
- · equilibrium can be attained both in homogeneous sheterogenous system.

# # Tupes of reaction based on direct

#### Reversible okn

- e exn take place in both direction
- · Chenerally these skn are take blace in close container
- · Reactant are never completely converted into the product.
- · Product can be converted into reactant.

# Isreversible skn

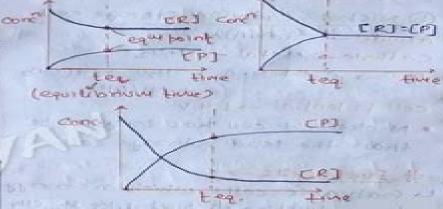
- o only in one direction.
- · in open container.
- · completely converted into product.
- · product can't be converted into reactant.

## Equilibrium state

1) Rate of horward reaction equal to reate of backward rkn.



constant at equilibrium etate.



- 3) Equilibrium state doesn't change by the presence of cotalyst.
- (4) catalyst increase the rate of borward & backward ste equally.
- (c) catalyst reduced the time to attained the equilibrium.
- @ Equilibrium dunamic in nature.
- A) The value of gibbs (AG) is zero at equilibrium etate.

# # Law of MORE action [LONA] ->

- Rate of chemical orkn is proportional to product of the active wase of the reactant is called LONA.

# **Chemical Equilibrium Solution Manual**

Thomas Rabe, Klaus Diedrich, Benno Runnebaum

#### **Chemical Equilibrium Solution Manual:**

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