

# Bond Energy

What makes a reaction endothermic or exothermic?

## Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** ( $\Delta H$ ). In this activity we will look at one way energy changes can be approximated for chemical reactions.

## Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )		Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
  - When bonds are (broken/formed) there is a positive energy change.
  - Breaking bonds is (endothermic/exothermic).
  - When bonds are (broken/formed) there is a negative energy change.
  - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.

- How do the changes in energy for the reverse reactions compare?

*The changes in energy for these reactions are the same, but opposite in sign.*

- Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

*If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.*

# Chemical Bonding Pogil Answers

**R. K. Sharma**



## **Chemical Bonding Pogil Answers:**

*Process Oriented Guided Inquiry Learning (POGIL)* Richard Samuel Moog, 2008 POGIL is a student centered group learning pedagogy based on current learning theory This volume describes POGIL s theoretical basis its implementations in diverse environments and evaluation of student outcomes *Chemical Bonds* Phillip Manning, 2009 Modern life is made up of a mind boggling array of materials A simple drinking cup for example might be made of Styrofoam paper or glass depending on the drinkers needs at the moment Home storage cabinets can be made of metal wood or plastic Space shuttles are assembled from silicon steel and hundreds of other materials All of these items owe their properties to the chemical bonds between the atoms that make up the substance Chemical Bonds examines the nature of the chemical bonds answering fundamental questions about how they form how they are broken and how they help define life as we know it **The Nature of the Chemical Bond and the Structure of Molecules and Crystals** Linus Pauling, 1960 Thorough discussion of the various types of bonds their relative natures and the structure of molecules and crystals **Chemical Bonding and the Geometry of Molecules** George E. Ryschkewitsch, 1963 [Chemical Bonding](#) M.S. Sethi & M. Satake, 2010 Contents Chemical Bonding I Basic Concepts Chemical Bonding II Additional Aspects Intermolecular Force and Crystal Structures

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**Chemical Bonds** Jeremy K. Burdett, 1997-05-28 Inorganic Chemistry This series reflects the breadth of modern research

in inorganic chemistry and fulfils the need for advanced texts The series covers the whole range of inorganic and physical chemistry solid state chemistry coordination chemistry main group chemistry and bioinorganic chemistry

**Chemical Bonds A Dialog** Jeremy K Burdett The University of Chicago USA Understanding the nature of the chemical bond is the key to understanding all chemistry be it inorganic physical organic or biochemistry In the form of a question and answer tutorial the fundamental concepts of chemical bonding are explored These range from the nature of the chemical bond via the regular hexagonal structure of benzene and the meaning of the term metallic bond to d orbital involvement in hypervalent compounds and the structure of  $N_2O$

**Chemical Bonds A Dialog** provides a novel format in terms of a dialog between two scientists insights into many key questions concerning chemical bonds an orbital approach to quantum chemistry

**The Chemical Bond** Gernot Frenking, Sason Shaik, 2014-06-13 This is the perfect complement to *Chemical Bonding Across the Periodic Table* by the same editors who are two of the top scientists working on this topic each with extensive experience and important connections within the community The resulting book is a unique overview of the different approaches used for describing a chemical bond including molecular orbital based valence bond based ELF AIM and density functional based methods It takes into account the many developments that have taken place in the field over the past few decades due to the rapid advances in quantum chemical models and faster computers

**Chemical Bonding Clarified Through Quantum Mechanics** George C. Pimentel, Richard D. Spratley, 1969 *Atomic Structure and Chemical Bonding, a Non-mathematical Introduction* Fritz Seel, 1963

**Chemistry of Chemical Bonding** R. K. Sharma, 2007

**Chemical Bonding in Solids and Fluids** Marcus Frederick Charles Ladd, 1994 Written by a highly regarded scientist and teacher this book examines and discusses the nature of and properties associated with interatomic and intermolecular forces in solids and fluids All discussions feature mathematical treatments accessible to chemistry students Provides stereoscopic diagrams of three dimensional structure covers computer methods where appropriate classifies and discusses solids in terms of bond type considers liquid structure and properties For students in chemistry chemical physics and biochemistry

**The Concept of the Chemical Bond** Dieter Cremer, Z.B. Maksic, 1990-06-13 The state of the art in contemporary theoretical chemistry is presented in this 4 volume set with numerous contributions from the most highly regarded experts in their field It provides a concise introduction and critical evaluation of theoretical approaches in relation to experimental evidence

*Chemical Bonding* Audrey L. Companion, 1979

*The Chemical Bond* Gernot Frenking, Sason Shaik, 2014-06-13 A unique overview of the different kinds of chemical bonds that can be found in the periodic table from the main group elements to transition elements lanthanides and actinides It takes into account the many developments that have taken place in the field over the past few decades due to the rapid advances in quantum chemical models and faster computers This is the perfect complement to *Chemical Bonding Fundamentals and Models* by the same editors who are two of the top scientists working on this topic each with extensive experience and important connections within the community

**Teaching Chemical**

**Bonding** Margaret Irene Lindsay, 1995 This document presents an instructional strategy for teaching chemical bonding using parables and music Games student interactions and worksheets are included in the lesson plans Topics include metallic bonding covalent bonding including molecular and network structure and ionic bonding JRH      **Structure and Bonding**  
Jack Barrett, 2001 This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes      **Chemical Bonding** Kok Koon Leong, 2021-05-03  
This book introduces the principles behind chemical bonding to teenagers between the ages of fifteen to seventeen Topics covered include ionic bonding covalent bonding and metallic bonding      **Chemical Bonds and Bond Energy** Robert Thomas Sanderson, 1971

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## **Table of Contents Chemical Bonding Pogil Answers**

1. Understanding the eBook Chemical Bonding Pogil Answers
  - The Rise of Digital Reading Chemical Bonding Pogil Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Bonding Pogil Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Bonding Pogil Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Bonding Pogil Answers
  - Personalized Recommendations
  - Chemical Bonding Pogil Answers User Reviews and Ratings
  - Chemical Bonding Pogil Answers and Bestseller Lists

5. Accessing Chemical Bonding Pogil Answers Free and Paid eBooks
  - Chemical Bonding Pogil Answers Public Domain eBooks
  - Chemical Bonding Pogil Answers eBook Subscription Services
  - Chemical Bonding Pogil Answers Budget-Friendly Options
6. Navigating Chemical Bonding Pogil Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Bonding Pogil Answers Compatibility with Devices
  - Chemical Bonding Pogil Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Bonding Pogil Answers
  - Highlighting and Note-Taking Chemical Bonding Pogil Answers
  - Interactive Elements Chemical Bonding Pogil Answers
8. Staying Engaged with Chemical Bonding Pogil Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Bonding Pogil Answers
9. Balancing eBooks and Physical Books Chemical Bonding Pogil Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Bonding Pogil Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemical Bonding Pogil Answers
  - Setting Reading Goals Chemical Bonding Pogil Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Bonding Pogil Answers
  - Fact-Checking eBook Content of Chemical Bonding Pogil Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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