

Types of Chemical Reactions Worksheet #1

H/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ **Synthesis**
- 2) $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$ **Single Replacement**
- 3) $3\text{KOH} + \text{AlCl}_3 \rightarrow \text{Al(OH)}_3 + 3\text{KCl}$ **Double Replacement**
- 4) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ **Synthesis**
- *5) $\text{Ca(OH)}_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$ **Double Replacement**
- 6) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ **Decomposition**
- 7) $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$ **Single Replacement**
- 8) $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$ **Double Replacement**
- 9) $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$ **Synthesis**
- *10) $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ **Single Replacement**

*Remember: "H₂O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement: $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement: $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition: $\text{Mg(ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement: $3\text{Ca(OH)}_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe(OH)}_3 + 3\text{CaCl}_2$
- 16) Single replacement: $\text{Fe} + \text{Cu(NO}_3)_2 \rightarrow \text{Fe(NO}_3)_2 + \text{Cu}$ [Assume Fe²⁺]
- 17) Decomposition: $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis: $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement: $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement: $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$ [Copper (II) is used here]

Chemical Reactions Guided Answers

Hussin A.Rothana



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A Pronouncing Chemical Formula Speller and Contest Guide Carl Alfred Jacobson, 1925

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J.

Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

Student Guide For Living Chemistry Barbara Ucko, 2012-12-02 Student Guide for Living Chemistry is a 23-chapter textbook guide that allows students to study and review on their own and test their understanding to help them prepare for examinations. Every chapter begins with a list of objectives stating exactly the skills to develop in a particular unit. Each objective corresponds to a section in the textbook Living Chemistry. Three kinds of questions are provided for each objective to check the student's understanding: namely, short answer, Study Questions, multiple choice, and fill-in. The answers for all questions are provided at the end of the chapter. The opening chapters cover the SI units, composition of matter, chemical bonding, compounds, chemical change, gases, respiration, and water. The subsequent chapters deal with solutions, acids, bases, salts, nuclear and organic chemistry, oxygen derivatives, and hydrocarbons, polymers, and other organic derivatives. This textbook also explores the chemistry of carbohydrates, lipids, proteins, enzymes, and energy and carbohydrate metabolism. The remaining chapters discuss the chemistry of vitamins, hormones, body fluids, drugs, and poisons. Undergraduate chemistry students will find this book invaluable.

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Strachan,2016-06-23 This edition of our successful series to support the Cambridge IGCSE Chemistry syllabus 0620 is fully updated for the revised syllabus from first examination from 2016 The Cambridge IGCSE Chemistry Practical Teacher s Guide complements the Practical Workbook helping teachers to include more practical work in lessons Specific support is provided for each of the carefully designed investigations to save teachers time The Teacher s Guide contains advice about planning investigations guidance about safety considerations differentiated learning suggestions to support students who might be struggling and to stretch the students who are most able as well as answers to all the questions in the Workbook The Teacher s Guide also includes a CD ROM containing model data to be used in instances when an investigation cannot be carried out *Survival Guide to General Chemistry* Patrick E. McMahon,Rosemary McMahon,Bohdan

Khomtchouk,2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields **Chemistry, Student Study**

Guide James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH₄ from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules

that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Study Guide for Psychology David G. Myers, Richard O. Straub, 2009-04-03 Longtime Myers collaborator Richard Straub provides an updated study guide for the new edition

[The Instructional Leader's Guide to Implementing K-8 Science Practices](#) Rebecca Lowenhaupt, Katherine L. McNeill, Rebecca Katsh-Singer, Benjamin R. Lowell, Kevin Cherbow, 2021-10-25 An accessible engaging primer on the eight science practices at the heart of the Next Generation Science Standards NGSS providing K 8 instructional leaders with the grounding they need to ensure excellent science instruction in every classroom The NGSS reconceptualize science instruction by redefining the teacher as someone who helps students construct their own knowledge by thinking like scientists and engaging in discrete science practices However with STEM teachers in short supply and generalists often feeling underprepared to teach elementary and middle school science what can instructional leaders do to ensure students get a strong start in this critical area and learn to love science Although a content neutral approach to supervision one that emphasizes general pedagogical features such as student engagement cognitive load or classroom management is undoubtedly beneficial the best instructional leaders know that content specific approaches are necessary to achieve real excellence We therefore need to go deeper if we want to engage both teachers and students with the science practices We need science specific supervision With that in mind the authors provide vignettes and examples of the science practices in use advice on observing science classrooms concrete look fors and guidance on fostering ongoing teacher learning They also offer a rich compendium of research and evidence based resources including sample lessons FAQs and more than a dozen downloadable tools to facilitate classroom observation feedback sessions and professional development This is an essential guide for any K 8 instructional leader who wants to empower all teachers to provide all students with rich science experiences and develop the cognitive and noncognitive skills students will need to thrive in more advanced courses work and society

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