



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes. In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time.</p> <p>Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.</p>	<p>Equilibrium Constant</p> <p>The equilibrium constant, K, is a measure of the extent to which a reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Graphical Representation of Equilibrium</p>												
<p>Equilibrium Constant, K_c</p> <p>The equilibrium constant, K_c, is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_p</p> <p>The equilibrium constant, K_p, is defined as the ratio of the partial pressures of the products to the partial pressures of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	<p>Equilibrium Constant, K_c vs. K_p</p> <p>The relationship between K_c and K_p is given by the equation:</p> $K_p = K_c (RT)^{\Delta n}$ <p>where R is the gas constant, T is the temperature in Kelvin, and Δn is the change in the number of moles of gas.</p>												
<p>Equilibrium Constant, K_a</p> <p>The equilibrium constant, K_a, is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_a = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_b</p> <p>The equilibrium constant, K_b, is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_b = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant, K_a vs. K_b</p> <p>The relationship between K_a and K_b is given by the equation:</p> $K_a K_b = K_w$ <p>where K_w is the ion-product constant for water, which is equal to 1.0×10^{-14} at 25°C.</p>												
<p>Le Chatelier's Principle: the effect of changing conditions on equilibrium</p> <table border="1"> <thead> <tr> <th>Condition</th> <th>Effect on Equilibrium</th> <th>Example</th> </tr> </thead> <tbody> <tr> <td>Concentration</td> <td> <ul style="list-style-type: none"> Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products. Increasing the concentration of a product will shift the equilibrium to the left, favoring the reactants. </td> <td> <p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the concentration of N_2 will shift the equilibrium to the right, favoring the formation of NH_3.</p> </td> </tr> <tr> <td>Pressure</td> <td> <ul style="list-style-type: none"> Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas. Decreasing the pressure will shift the equilibrium to the side with the most moles of gas. </td> <td> <p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the pressure will shift the equilibrium to the right, favoring the formation of NH_3.</p> </td> </tr> <tr> <td>Temperature</td> <td> <ul style="list-style-type: none"> Increasing the temperature will shift the equilibrium to the side that is endothermic. Decreasing the temperature will shift the equilibrium to the side that is exothermic. </td> <td> <p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the temperature will shift the equilibrium to the left, favoring the reactants.</p> </td> </tr> </tbody> </table>			Condition	Effect on Equilibrium	Example	Concentration	<ul style="list-style-type: none"> Increasing the concentration of a reactant will shift the equilibrium to the right, favoring the products. Increasing the concentration of a product will shift the equilibrium to the left, favoring the reactants. 	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the concentration of N_2 will shift the equilibrium to the right, favoring the formation of NH_3.</p>	Pressure	<ul style="list-style-type: none"> Increasing the pressure will shift the equilibrium to the side with the fewest moles of gas. Decreasing the pressure will shift the equilibrium to the side with the most moles of gas. 	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the pressure will shift the equilibrium to the right, favoring the formation of NH_3.</p>	Temperature	<ul style="list-style-type: none"> Increasing the temperature will shift the equilibrium to the side that is endothermic. Decreasing the temperature will shift the equilibrium to the side that is exothermic. 	<p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the temperature will shift the equilibrium to the left, favoring the reactants.</p>
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Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter, Steven S. Zumdahl, 2004-04

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex

macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *Resources in Education* ,1996-05

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Study Guide for Chemical Principles, Fourth Edition, by Dickerson, Gray, Darensbourg, and Darensbourg Patricia L. Samuel, 1984

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