



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

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Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time having reached Equilibrium. Equilibrium reactions are represented by the symbol \rightleftharpoons .

<p>What is Equilibrium?</p> <p>Equilibrium is a state of balance between two or more opposing processes.</p> <p>In chemistry, it refers to a state where the concentrations of reactants and products remain constant over time.</p> <p>Equilibrium is reached when the rate of the forward reaction is equal to the rate of the reverse reaction.</p> <p>Equilibrium is a dynamic process, meaning that the reactions are still occurring, but the concentrations of the species remain constant.</p>	<p>Equilibrium Constant</p> <p>The equilibrium constant (K) is a measure of the extent to which a chemical reaction proceeds. It is defined as the ratio of the concentrations of the products to the concentrations of the reactants, each raised to the power of its stoichiometric coefficient.</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Graphical Representation of Equilibrium</p> <p>The graph shows the concentration of reactants and products over time. The reactant concentration starts high and decreases exponentially towards a constant value. The product concentration starts at zero and increases exponentially towards a constant value. The equilibrium is reached when the concentrations of reactants and products remain constant over time.</p>
<p>Equilibrium Constant, K_c</p> <p>The equilibrium constant K_c is calculated using the concentrations of the species in the reaction mixture.</p> <p>For the reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K_c is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ <p>where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the species in mol/L.</p>	<p>Equilibrium Constant, K_p</p> <p>The equilibrium constant K_p is calculated using the partial pressures of the gases in the reaction mixture.</p> <p>For the reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K_p is given by:</p> $K_p = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$ <p>where P_A, P_B, P_C, and P_D are the partial pressures of the gases in atm.</p>	
<p>Equilibrium Constant, K</p> <p>The equilibrium constant K is a general term that can refer to either K_c or K_p, depending on the context. It is a measure of the extent to which a chemical reaction proceeds.</p> <p>For the reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant K is given by:</p> $K = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ <p>or</p> $K = \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b}$	<p>Relationship between K_c and K_p</p> <p>The relationship between K_c and K_p is given by the equation:</p> $K_p = K_c (RT)^{\Delta n}$ <p>where R is the gas constant (8.314 J/mol·K), T is the temperature in Kelvin, and Δn is the change in the number of moles of gas.</p>	
<p>Le Chatelier's Principle: the effect of changing conditions on equilibrium</p>		
<p>Concentration</p>	<p>Effect on Equilibrium</p> <ul style="list-style-type: none">Increasing the concentration of a reactant shifts the equilibrium to the right.Increasing the concentration of a product shifts the equilibrium to the left.	<p>Example</p> <p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the concentration of N_2 shifts the equilibrium to the right, producing more NH_3.</p>
<p>Pressure</p>	<p>Effect on Equilibrium</p> <ul style="list-style-type: none">Increasing the pressure shifts the equilibrium to the side with fewer moles of gas.Decreasing the pressure shifts the equilibrium to the side with more moles of gas.	<p>Example</p> <p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the pressure shifts the equilibrium to the right, producing more NH_3.</p>
<p>Temperature</p>	<p>Effect on Equilibrium</p> <ul style="list-style-type: none">Increasing the temperature shifts the equilibrium to the endothermic side.Decreasing the temperature shifts the equilibrium to the exothermic side.	<p>Example</p> <p>For the reaction:</p> $N_2 + 3H_2 \rightleftharpoons 2NH_3$ <p>Increasing the temperature shifts the equilibrium to the left, producing less NH_3.</p>

Chemical Equilibrium Study Guide Answer Sheet

Mr. Rohit Manglik



Chemical Equilibrium Study Guide Answer Sheet:

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