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Chemical Equilibrium

1. For the reversible reaction,

$$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g}) + \text{heat}$$
the equilibrium shifts in forward direction [2014]

- (a) by increasing the concentration of $\text{NH}_3(\text{g})$
(b) by decreasing the pressure
(c) by decreasing the concentrations of $\text{N}_2(\text{g})$ and $\text{H}_2(\text{g})$
(d) by increasing pressure and decreasing temperature
2. Using the Gibbs energy change, $\Delta G^\circ = +63.3 \text{ kJ}$ for the following reaction,
$$\text{Ag}_2\text{CO}_3(\text{s}) \rightleftharpoons 2\text{Ag}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq})$$
the K_{sp} of $\text{Ag}_2\text{CO}_3(\text{s})$ in water at 25°C is ($R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$) [2014]
- (a) 3.2×10^{-26}
(b) 8.0×10^{-12}
(c) 2.9×10^{-3}
(d) 7.9×10^{-7}

3. KMnO_4 can be prepared from K_2MnO_4 as per reaction,



The reaction can go to completion by removing OH^- ions by adding [2013]

- (a) HCl
(b) KOH
(c) CO_2
(d) SO_2

4. The value of ΔH for the reaction,



is less than zero. Formation of $\text{XY}_4(\text{g})$ will be favoured at [2011]

- (a) low pressure and low temperature
(b) high temperature and low pressure
(c) high pressure and low temperature
(d) high temperature and high pressure
5. For the reaction,
$$\text{N}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}(\text{g})$$
, the equilibrium constant is K_1 . The equilibrium constant is K_2 for the reaction,
$$2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$$
. What is K for the reaction, $\text{NO}_2(\text{g}) \rightleftharpoons \frac{1}{2}\text{N}_2(\text{g}) + \text{O}_2(\text{g})$? [2011]
- (a) $1/(4K_1K_2)$
(b) $[1/(K_1K_2)]^{1/2}$
(c) $1/(K_1K_2)$
(d) $1/2K_1K_2$
6. In which of the following equilibrium K_c and K_p are not equal? [2010]

- (a) $2\text{NO}(\text{g}) \rightleftharpoons \text{N}_2(\text{g}) + \text{O}_2(\text{g})$
(b) $\text{SO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons \text{SO}_3(\text{g}) + \text{NO}(\text{g})$
(c) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
(d) $2\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{CO}_2(\text{g})$

7. The dissociation constants for acetic acid and HCN at 25°C are 1.5×10^{-5} and 4.5×10^{-10} , respectively. The equilibrium constant for the equilibrium,



- (a) 3.0×10^5
(b) 3.0×10^{-5}
(c) 3.0×10^{-4}
(d) 3.0×10^4

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**John Krenos, Joseph Potenza, Carl
Hoeger**



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