

## Chemical Bonding Worksheet

Ionic Bond  
 Covalent Bond  
 Metallic Bond

between a Metal and Non-Metal  
 between a Non-Metal and Non-Metal  
 between a Metal and Metal

Compound	Name	Oxidation states of each atom	Bond Type
NO <sub>2</sub>	nitrogen dioxide	N = +4 O = -2	covalent
NaCl	sodium chloride	Na = +1 Cl = -1	ionic
SO <sub>2</sub>	sulfur dioxide	S = +4 O = -2	covalent
PO <sub>4</sub> <sup>3-</sup>	phosphate anion	P = +5 O = -2	covalent
MgBr <sub>2</sub>	magnesium bromide	Mg = +2 Br = -1	ionic
CaO	calcium oxide	Ca = +2 O = -2	ionic
H <sub>2</sub> O	water	H = +1 O = -2	covalent
K <sub>2</sub> O	potassium oxide	K = +1 O = -2	ionic
Cu-Zn alloy	brass	Cu = 0 Zn = 0	metallic
O <sub>2</sub>	oxygen gas	O = 0	covalent
CuCl <sub>2</sub>	copper(II) chloride	Cu = +2 Cl = -1	ionic
NO <sub>2</sub> <sup>-</sup>	nitrite anion	N = +3 O = -2	covalent
TiO <sub>2</sub>	titanium (IV) oxide	Ti = +4 O = -2	ionic
HF	hydrofluoric acid	H = +1 F = -1	covalent
Rb <sub>2</sub> S	rubidium sulfide	Rb = +1 S = -2	ionic
Au-Pt mixture	white gold	Au = 0 Pt = 0	metallic
Fe <sub>2</sub> O <sub>3</sub>	iron (III) oxide	Fe = +3 O = -2	ionic
C <sub>6</sub> H <sub>12</sub> O <sub>6</sub>	glucose	C = +4 H = -1 O = -2	covalent

# Chemical Bonds Guide Answers

**Peipei Pang**



## Chemical Bonds Guide Answers:

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