

NUCLEAR CHEMISTRY STUDY GUIDE—CH. 21

1. Write the symbol for uranium—238 in isotope format (with top and bottom #s).
2. What does the top # represent?
3. What does the bottom # represent?
4. What particles are found in the nucleus?
5. Why does the strong nuclear force exist?
6. What is the electrostatic force?
7. What particles does the electrostatic force deal with?
8. Define nuclear reactions.
9. Define nuclear fission and fusion.
10. Nuclear reactions usually produce _____ amounts of energy, and isotopes become _____ stable. The _____ number and the _____ number may also change.
11. What is transmutation?
12. Write the beta decay of radium—226.
13. Write the alpha decay of radium—226.
14. Why would the isotopes and the end of problems #13 and #14 still be unstable?
15. Define nuclear radiation.
16. List the 3 radioactive particles in order of INCREASING penetrating power.
17. What is radioactive decay?
18. Alpha particles are basically _____ nuclei, and beta particles are basically high energy _____.
19. What is half life?
20. What is radioactive dating?
21. What is cobalt—60 used for?
22. Does fission or fusion produce nuclei with heavier mass than the original mass?
23. Name 3 differences between chemical and nuclear reactions.
24. What is carbon—13 radioactive?
25. Elements past atomic # _____ are radioactive.
26. If $1/16$ of a sample of a radioactive isotope is left after nuclear decay, how long did this take if the $\frac{1}{2}$ life is 4 days?
27. If the half life of a radioactive isotope is 4 days, how much of a 50 gram sample is left after 12 days?

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Table of Contents Chemhnhs Pearson Study Guide Nuclear Chemistry

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 - The Rise of Digital Reading Chemhnhs Pearson Study Guide Nuclear Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemhnhs Pearson Study Guide Nuclear Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemhnhs Pearson Study Guide Nuclear Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemhnhs Pearson Study Guide Nuclear Chemistry
 - Personalized Recommendations
 - Chemhnhs Pearson Study Guide Nuclear Chemistry User Reviews and Ratings

- Chemhnhs Pearson Study Guide Nuclear Chemistry and Bestseller Lists
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 - Participating in Virtual Book Clubs
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 - Benefits of a Digital Library
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- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemhnhs Pearson Study Guide Nuclear Chemistry
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 - Carving Out Dedicated Reading Time
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 - Interactive and Gamified eBooks

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