

# Precipitation Reactions

- Reactions that produce insoluble products (or precipitates) when two aqueous solutions are mixed.
- Examples:
- (1)  $\text{AgNO}_3(aq) + \text{NaCl}(aq) \rightarrow \text{AgCl}(s) + \text{NaNO}_3(aq)$
- (2)  $\text{Pb}(\text{NO}_3)_2(aq) + \text{K}_2\text{CrO}_4(aq) \rightarrow \text{PbCrO}_4(s) + 2\text{KNO}_3(aq)$
- (3)  $\text{BaCl}_2(aq) + \text{H}_2\text{SO}_4(aq) \rightarrow \text{BaSO}_4(s) + 2\text{HCl}(aq)$
- (4)  $3\text{AgNO}_3(aq) + \text{Na}_3\text{PO}_4(aq) \rightarrow \text{Ag}_3\text{PO}_4(s) + 3\text{NaNO}_3(aq)$

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**E Durkheim**



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**Inquiry-based Experiments in Chemistry** Valerie Ludwig Lechtanski, 2000 Inquiry Based Experiments in Chemistry is an alternative to those cookbook style lab manuals providing a more accurate and realistic experience of scientific investigation and thought for the high school chemistry or physical science student

**The Directory of U.S. Trademarks**, 1993

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**General Chemistry Key Concise Review Guide Sixth Edition** Darrell D. Ebbing, 1999

*Chemistry and Student Guide and Lab Experiments and Chemistry on Internet Package* Theodore E. Brown, H. Eugene LeMay, Bruce E. Bursten, 1998-05

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