

Chemical Formula Writing Worksheet Solutions

Write chemical formulas for the compounds in each box. The names are found by finding the intersection between the cations and anions. Example: The first box is the intersection between the "zinc" cation and the "chloride" anion, so you should write " ZnCl_2 ", as shown.

| | <i>zinc</i> | <i>iron (II)</i> | <i>iron (III)</i> | <i>gallium</i> | <i>silver</i> | <i>lead (IV)</i> |
|-----------------|---|---|---|---|--|---|
| <i>chloride</i> | ZnCl_2 | FeCl_2 | FeCl_3 | GaCl_3 | AgCl | PbCl_4 |
| <i>acetate</i> | $\text{Zn}(\text{C}_2\text{H}_3\text{O}_2)_2$ | $\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_2$ | $\text{Fe}(\text{C}_2\text{H}_3\text{O}_2)_3$ | $\text{Ga}(\text{C}_2\text{H}_3\text{O}_2)_3$ | $\text{Ag C}_2\text{H}_3\text{O}_2$ | $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_4$ |
| <i>nitrate</i> | $\text{Zn}(\text{NO}_3)_2$ | $\text{Fe}(\text{NO}_3)_2$ | $\text{Fe}(\text{NO}_3)_3$ | $\text{Ga}(\text{NO}_3)_3$ | AgNO_3 | $\text{Pb}(\text{NO}_3)_4$ |
| <i>oxide</i> | ZnO | FeO | Fe_2O_3 | Ga_2O_3 | Ag_2O | PbO_2 |
| <i>nitride</i> | Zn_3N_2 | Fe_3N_2 | FeN | GaN | Ag_3N | Pb_3N_4 |
| <i>sulfate</i> | ZnSO_4 | FeSO_4 | $\text{Fe}_2(\text{SO}_4)_3$ | $\text{Ga}_2(\text{SO}_4)_3$ | Ag_2SO_4 | $\text{Pb}(\text{SO}_4)_2$ |

Write the formulas for the following compounds:

- 1) copper (II) chloride **CuCl_2**
- 2) lithium acetate **$\text{LiC}_2\text{H}_3\text{O}_2$**
- 3) vanadium (III) selenide **V_2Se_3**
- 4) manganese (IV) nitride **Mn_3N_4**
- 5) beryllium oxide **BeO**
- 6) sodium sulfate **Na_2SO_4**
- 7) aluminum arsenide **AlAs**
- 8) potassium permanganate **KMnO_4**
- 9) chromium (VI) cyanide **$\text{Cr}(\text{CN})_6$**
- 10) tin (II) sulfite **SnSO_3**
- 11) vanadium (V) fluoride **VF_5**
- 12) ammonium nitrate **NH_4NO_3**

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Rachel Turoscy



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