

Chemistry 33

The Ideal Gas Law

(Review for Chemistry 33)

Information:

- **Boyle's Law (1662)** – Volume or Molecular Occupancy – $VT(C) = \text{const.}$
 $P(C) \text{ always } \propto \frac{1}{V}$
- **Charles Law (1828)** – The volume of a sample of a gas varies directly with pressure, if the temperature is held constant.

$$V = k_1 \frac{1}{P} \quad \text{if constant is } k_1 T$$
 where k_1 is for the number of moles of gas.
- **Charles Law (1828)** – The volume of a gas varies directly with temperature, if the pressure is held constant.

$$V = k_2 T \quad \text{if constant is } k_2 P$$
- **Avogadro's Hypothesis (1811)** – Samples of different gases, which contain the same number of molecules – of any complexity, size, or shape – occupy the same volume at the same temperature and pressure.

$$V = k_3 n \quad \text{if constant is } k_3 P \text{ and } T$$

Model 1: The Ideal Gas Law Equation.

$$PV = nRT$$

where R is a constant called the ideal gas constant.

The numerical value of the ideal gas constant is calculated from the fact that one mole of gas occupies 22.414 L at a pressure of one atmosphere and a temperature of 0°C (273.15 K):

$$R = \frac{P}{n} = \frac{1 \text{ atm} (101.325 \text{ kPa})}{1 \text{ mole} (6.022 \times 10^{23})} = \text{about } \frac{1 \text{ atm}}{6.022 \times 10^{23}}$$

Chemactivity 33 The Ideal Gas Law

Ensheng Dong



Chemactivity 33 The Ideal Gas Law:

Unveiling the Power of Verbal Beauty: An Psychological Sojourn through **Chemactivity 33 The Ideal Gas Law**

In a world inundated with screens and the cacophony of quick transmission, the profound power and mental resonance of verbal art usually disappear in to obscurity, eclipsed by the regular barrage of noise and distractions. Yet, set within the lyrical pages of **Chemactivity 33 The Ideal Gas Law**, a charming perform of literary brilliance that impulses with organic thoughts, lies an unforgettable trip waiting to be embarked upon. Written by way of a virtuoso wordsmith, this exciting opus manuals viewers on an emotional odyssey, lightly revealing the latent potential and profound impact stuck within the complex internet of language. Within the heart-wrenching expanse of this evocative analysis, we can embark upon an introspective exploration of the book is key subjects, dissect its interesting publishing style, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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