

CHEM 108
ChemActivity 23: Intermolecular Forces
Solution Key

The Model: Dipole-Dipole Forces

Within a polar molecule there is a separation of charge. Neighboring molecules in a bulk sample of a polar compound will arrange themselves in order to maximize attractive forces and to minimize repulsive forces between the molecules. (In a liquid there is still enough thermal energy that the molecules are constantly moving past each other; so this is a "time-averaged" picture we're considering.) The net attractive forces between polar molecules are called **dipole-dipole forces**, which are typically only about 1% as strong (~ 4 kJ/mol) as a covalent bond.



Key Questions

1. When a substance boils, at the surface of the liquid inter-particle attractive forces must be broken as particles go from the liquid into the corresponding vapor phase.



Coulomb's Law
$$F = k \frac{Q_+ Q_-}{r^2}$$



- a. The attractive forces between pairs of particles in a polar molecular substance (on the left, in which $\delta+$ represents a partial positive charge and $\delta-$ represents a partial negative charge) is weaker stronger than that in an ionic substance (on the right). (Circle your answer.)
- b. As attractive forces between neighboring molecules in a bulk sample get stronger, it will require less more heat energy to separate molecules from each other. (Circle your answer.)
- c. According to Coulomb's Law, the boiling point of molecular substances should be lower higher than the boiling point of ionic substances. (Circle your answer. Is this consistent with observation?)

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